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Citation for published version:

Nazeri Ghogh, M, Chapoy, A, Burgass, RW & Tohidi Kalorazi, B 2017, 'Measured densities and derived thermodynamic properties of CO₂-rich mixtures in gas, liquid and supercritical phases from 273 K to 423 K and pressures up to 126 MPa', *Journal of Chemical Thermodynamics*, vol. 111, pp. 157-172.
<https://doi.org/10.1016/j.jct.2017.03.036>

Digital Object Identifier (DOI):

[10.1016/j.jct.2017.03.036](https://doi.org/10.1016/j.jct.2017.03.036)

Link:

[Link to publication record in Heriot-Watt Research Portal](#)

Document Version:

Peer reviewed version

Published In:

Journal of Chemical Thermodynamics

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Measured densities and derived thermodynamic properties of CO₂-rich mixtures in gas, liquid and supercritical phases from 273 K to 423 K and pressures up to 126 MPa

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ABSTRACT

The densities of three multi-component mixtures with high CO₂ content and common impurities, i.e. hydrocarbons, nitrogen, hydrogen, oxygen, argon and carbon monoxide, were measured using an Anton Paar DMA-HPM densitometer. The mixtures include MIX 1 with 0.9564 mole fraction CO₂ and 0.0436 mole fraction impurities of methane and non-condensable gases, MIX 2 with 0.8983 mole fraction CO₂ and 0.1017 mole fraction impurity of non-condensable gases and MIX 3 with 0.6999 mole fraction of CO₂ and 0.3001 mole fraction of light hydrocarbons. First, the densitometer was calibrated using pure CO₂ and then the density measurements of mixtures were carried out in the gas, liquid and supercritical phases at pressures up to 126 MPa at various isotherms of $T/K = 273, 283, 298, 323, 373$ and 423. The obtained data then were employed to evaluate the classical cubic equations of state (SRK and PR). In addition, the CO₂ volume correction term and the Peneloux shift parameter were introduced to improve density predictions. Comparisons show that applying CO₂ volume correction term to SRK EoS with modified k_{ij} could improve the density predictions and the AAD was reduced from 4.7% to 1.9%. In addition to the classical cubic equations of state, the new measured data were used to evaluate the GERG EoS. The AAD of the GERG EoS from experimental data measured in this work are 2.8%, 1.0% and 1.2% in the gas, liquid and supercritical phases, respectively. The overall AAD of GERG EoS from the new density data in this work is 1.7%. Thermodynamic properties, i.e. compressibility factor, specific heat capacity, speed of sound and Joule-Thomson coefficient, were then calculated from the

experimental density data using thermodynamic equations. Following this the derived properties were compared to the predictions made with the GERG equation of state.

Keywords: Density measurement, high pressure, Carbon capture transport and storage (CCS), CO₂ and impurities, Thermodynamic properties, effect of impurities

1. Introduction

The application of carbon capture and storage (CCS) has become increasingly important, from both scientific and industrial points of view, the overall aim being to reduce CO₂ emissions. CCS is a promising technology-based solution to reduce the significant amount of CO₂ release from large-scale industrial sources. In the CCS full chain, the CO₂ emissions from the industry sector need to be captured with emerging technologies such as pre-combustion, post-combustion and oxyfuel processes. The captured CO₂ then can be transported either by pipelines or ships to storage locations such as depleted oil and gas reservoirs, saline aquifers or unmineable coal seams [1]. Transport by pipeline is the preferred option for the large quantities of carbon dioxide over longer distances [2]. Technically, CO₂ can be transported through pipelines in the form of a gas, a supercritical fluid or in the sub-cooled liquid state. Operationally, the most efficient CO₂ pipelines used for enhanced oil recovery transport the CO₂ as a supercritical fluid [3]. As the critical point of CO₂ ($P_c = 7.38$ MPa and $T_c = 304.3$ K) and triple point ($P_t = 0.518$ MPa at $T_t = 216.6$ K) are very different from conventional fluids present in transport pipelines in oil and gas industry, the modeling of these pipelines poses new challenges [4]. It was proposed that the operating pressure of CO₂ transport pipelines should be above 8.6 MPa to make sure that the fluid will always be in the single super critical phase over a range of temperatures that the pipeline may encounter [5]. The captured CO₂ in the first step may contain various types and levels of impurities depending on the source and capturing technology [6][7][8]. The presence of impurities, particularly hydrogen and nitrogen, can change the fluids thermophysical properties such as critical pressure, density and viscosity, which could compromise the performance of the CCS processes. In the design of CO₂ transport pipeline, the overestimation or underestimation of the fluid physical properties, particularly the density of CO₂ mixtures with impurities, may results in the underdesign or overdesign of the pipeline diameter. This can cause operational problems such as high velocity and erosion or phase change and causing multiphase flow in the pipeline [9][10][11]. CO₂ and CO₂-rich

mixtures also play an important role in CO₂ enhanced oil recovery. The CO₂-rich mixtures may cause hydrate formation problem during the transportation [12][13].

Equations of state (EoS) are imperative to predict and understand the thermodynamic properties of fluids. Several equations of state are available in literature to predict the properties of high CO₂ content mixtures [14][15][16][17]. In this work, new measured density data are presented for three multi component CO₂-rich mixtures. The classical cubic equations of state Soave-Redlich-Kwong (SRK-EoS) [18] and Peng-Robinson (PR-EoS) [19] with CO₂ volume correction [20][21] (SRK-CO₂ and PR-CO₂) and Peneloux shift parameter [22] (SRK-Pen and PR-Pen) as well as the GERG EoS then were evaluated using the measured density data. The thermodynamic properties, i.e. compressibility factor, specific heat capacity, speed of sound and joule-Thomson coefficient also were calculated from the measured density data and were compared with the GERG EoS [23].

2. Literature review

According to the requirements of engineering applications for design and operation of CO₂ capture and storage systems, cubic equation of states are preferable to predict VLE properties and density calculations [16] due to the simplicity and availability in the oil and gas industry as well as commercial software packages. The systems to be studied may contain a wide range of components including pure CO₂, and mixtures with other gases, amines, ionic liquids, water, and brines. Some studies were conducted to investigate thermodynamic properties of CO₂ and CO₂-mixture systems using equations of state.

The SRK EoS was investigated by Frey et al. [24] for density and phase equilibria of mixtures, including the CO₂-H₂O and CO₂-CH₄ binary systems. They applied a density and temperature dependent volume translation function on SRK EoS. They found that selection of mixing rules has a significant influence on the results of their method, which is abbreviated as DMT. Also, Thiery et al. [25] evaluated the SRK EoS for VLE and volume calculations of CO₂-N₂, CO₂-CH₄ and CO₂-CH₄-N₂. Their results showed that with the SRK EoS, the average deviation for the saturated pressures is around 1% in the temperature range of 208.45-270 K for the CO₂-CH₄ system, 4% in the temperature range of 218.15-273.15 K for the CO₂-N₂ system, and 2-3% for the CO₂-CH₄-N₂ system.

The PR and Patel-Teja (PT) EoS were investigated by Al-Sahhaf et al. [26] for VLE of the N₂-CO₂-CH₄ ternary system. Also, Boyle and Carroll [27] investigated PR, SRK, PT, PR-Peneloux, SRK-Peneloux and PR-Mathias EoS for density calculations of CO₂-H₂S. The results showed that PT is the most accurate EoS in liquid region, supercritical region, and

overall, with an AAD of 2.16%, 2.26% and 1.82% respectively; while SRK is the most accurate EoS in the vapor region with an AAD of 0.51%. Seven cubic equations of state were evaluated with respect to VLE and density of CO₂ mixtures including CH₄, N₂, O₂, H₂S, SO₂ and Ar using many of the experimental data presented by [Li et al. \[16\] \[28\]\[29\]](#). The EoS evaluated were PR, RK, SRK, PT, PR-Peneloux, SRK-Peneloux and the improved SRK. The binary interaction parameters, k_{ij} , were calibrated with respect to VLE data.

[Mantovani et al. \[30\]](#) presented experimental data for supercritical CO₂ binary systems of nitrogen, oxygen and argon used in oxy-fuel capture process with almost 5% and 10% impurities. They used vibrating tube densitometer for pressures ranging from 1 MPa to 20 MPa at different temperatures from 303 to 383 K. They also tuned the binary interaction parameters against experimental data using PR, SRK-Peneloux and Benedict-Webb-Rubin-Starling (BWRS) equations of state. For the CO₂-N₂ systems, they reported an AAD of 2.10%, 3.05% and 1.71% for PR, SRKP and BWRS, respectively. Also, the AAD of 2.37%, 3.92% and 1.97% for those of CO₂-O₂ systems and 2.56%, 4.07% and 1.75% for those of CO₂-Ar systems. All the AADs reported using the new regressed k_{ij} parameters. It can be seen that for each case, BWRS can predicts better than cubic equation of states. [Sanchez-Vicente et al. \[31\]](#) presented the density data for three CO₂-H₂ mixtures, as the main impurity of pre-combustion capturing process, with 2%, 7.5% and 10% of hydrogen at 288.15–333.15 K and pressures between 1.5 and 23 MPa. They then compared their density data with the values calculated by the GERG-2004 equation of state using the original parameters provided by [Kunz et al. \[15\]](#). The deviations between the experimental and calculated density are also calculated and analyzed in the critical and liquid regions of the mixtures. They have concluded that 2% hydrogen can reduce the molar density of CO₂ up to 25% in the critical region which can significantly affect the compression and transportation development for CCS. They also have found that the GERG-2004 EoS can accurately predict the density of CO₂-rich systems with low H₂ concentration (2%) with an AAD of 0.6% while for the high H₂ concentrations, AADs of up to 4% and 14% were observed for the liquid phase and supercritical phase, respectively.

[Rivas et al. \[32\]](#) measured density of CO₂-CH₄ and CO₂-CO systems at $T/K = 304.21$ and 308.14 and pressures ranging from 0.1 to 20 MPa and x_{CO_2} (mole fraction of carbon dioxide in vapor phase) > 0.97 . Then, they compared the volumetric behavior of these systems to the PR, PC-SAFT and GERG equations of state. The deviations were reported less than 3.5% for PR, 2.8% for rescaled PC-SAFT and 1.0% for GERG. The VLE experimental data for various binary systems containing CO₂ were measured at various temperature and pressures for CO₂-Ar [\[33\]](#), CO₂-H₂ [\[34\]](#), CO₂-N₂ [\[34\]\[35\]](#) and CO₂-O₂ [\[36\]](#). The density of 0.95 CO₂ – 0.05

Ar was measured by Yang et al. [37] in the vapor phase at pressures up to 9.0 MPa. The experimental data also were evaluated by GERG-2008 and EOS-CG equations of state with the relative deviation of 0.95% and 0.18%, respectively. The densities of 0.95 CO₂ – 0.05 H₂S also was measure by Nazeri et al. [21] at pressures up to 42 MPa and various isotherms in the gas, liquid and supercritical phases. The densities of two ternary systems of Ar-N₂-CO₂ with different concentrations were measured by Yang et al. [38] at pressures from 3 to 31 MPa at temperatures from 323 to 423 K in supercritical phase. The deviation of GERG-2008 from experimental data was reported to be 0.6%.

3. Experimental

3.1 Material

Varieties of multi-component mixtures with diverse impurities and different percentages were prepared, as shown on Table 1. Each mixture represents the composition of gas from a specific source. MIX 1 with 0.0436 mole fraction of impurities such as methane, nitrogen, hydrogen, argon, oxygen and carbon monoxide can show a proper behavior of the streams suitable for the carbon capture, transport and storage. MIX 2 with a wider range of 0.1017 mole fraction of non-condensable gases, i.e., nitrogen, oxygen and argon, was also prepared to mimic the CCS stream fluids. Mixture 3 with 0.3001 mole fraction of light hydrocarbons and 0.6999 mole fraction CO₂ represents the composition of a gas field in South East Asia. The mixtures supplied by BOC were certified on the basis of gravimetry in accordance with ISO 6142 with analytical validation. The volume of the cylinders were about 2 cubic meters. It is recommended by the supplier not to use the product below 5% of actual contents. According to the suppliers' instructions to prevent condensation, the cylinders were kept in the laboratory area with a temperature of about 293.15 K (20 °C). The reported expanded uncertainties in each table are based on a standard uncertainty multiplied by a coverage factor $k = 2$, providing a level of confidence of approximately 95%.

3.2 Equipment description

The densities of CO₂-rich systems were measured using a high temperature and pressure oscillating U-tube densitometer, Anton Paar DMA-HPM, which consists of a measuring cell and an evaluation unit. A schematic view of the apparatus is shown in Figure 1. The equipment was described in the previous publication [20]. Briefly, the measuring cell includes a U-shaped Hastelloy C-276 tube that is excited to vibrate at its characteristic frequency electronically. The

DMA-HPM is connected to an mPDS 2000V3 evaluation unit which measures the oscillation period. The resolution of the unit is seven significant digits.

The temperature of the densitometer is controlled by an oven, manufactured by BINDER GmbH, which can be used at temperatures between $-70\text{ }^{\circ}\text{C}$ to $200\text{ }^{\circ}\text{C}$ ($203 - 473\text{ K}$). The measuring U-shaped cell is insulated from the environment keeping the temperature stable to $\pm 0.02\text{ K}$. A built-in thermometer which is connected to the mPDS 2000V3 unit shows the temperature of the vibrating tube cell. A hand pump which can inject or withdraw the mercury to the set-up is used to control the system pressure. Two Quartzdyne pressure transducer (model: QS 30K-B) with the design pressure up to 207 MPa and standard uncertainty of $\pm 0.02\text{ MPa}$ [39] were connected to record the system pressure.

3.3 Measurement and calibration procedure

All the experiments were conducted using the Anton Paar densitometer. In each test, after applying vacuum to the entire system, the sample was injected through the injection point on top of the densitometer. It should be pointed out that the sample fluid during the injection was kept in single phase to avoid composition change due to flashing. Then after disconnecting the sample cylinder from the system, it was allowed to stabilize at the desired temperature. When the temperature of the vibrating tube was stable, the desired pressure was set using the hand pump. Once conditions had stabilized, the oscillation period of the U-tube was determined from the interface mPDS 2000V3 evaluation unit.

The measurement of density with a vibrating tube densitometer is not absolute, thus, the raw data (period of oscillation) should be further treated to obtain the densities. The relationship between them is:

$$\rho(T, P) = A(T, P)\tau^2(T, P) - B(T, P) \quad (1)$$

where $\rho(T, P)$ is the sample density at temperature T and pressure P , $\tau(T, P)$ is the period of oscillation at temperature and pressure, $A(T, P)$ and $B(T, P)$ are the apparatus parameters depending on temperature and pressure, and they must be determined from calibration measurements. For calibration, CO_2 density can be used as a reference substance at two different pressures (the lowest and the highest desired pressures in the system at the same temperature) in gas, liquid and supercritical phases. The apparatus parameters then were defined as follows:

$$A(T, P) = \frac{\rho(T, P_1) - \rho(T, P_2)}{\tau^2(T, P_1) - \tau^2(T, P_2)} \quad (2)$$

$$B(T,P) = \frac{\tau^2(T,P_2)\rho(T,P_1) - \tau^2(T,P_1)\rho(T,P_2)}{\tau^2(T,P_1) - \tau^2(T,P_2)} \quad (3)$$

During our calibration, the densities of pure CO₂ were measured at different desired pressures for each isotherm. The density data used for calibration were calculated with the Span and Wagner multi-parameter equation of state [40]. Then, the parameters *A* and *B* were calculated by plotting the linear trend line for density versus squared oscillation period measured at different desired pressures at each isotherm. The procedure was repeated for each isotherm once at low pressures, i.e. gas phase, then at higher pressures, i.e. dense liquid / supercritical phases. Figure 2 and Figure 3 show the procedure to determine *A* and *B* parameters at 373.15 K (100 °C) at low and high pressures, respectively. All calibration data for each isotherm at low pressures (gas phase) and high pressures (dense phase) can be seen in Table 2 and Table 3, respectively. Finally, the *A* and *B* calibration parameters at low and high pressures at different measuring temperatures can be found in Table 4.

3.4 Density validation

The measured densities for pure CO₂ and MIX 1 were compared to the density data measured by Al-Siyabi et al. [41]. As shown on Figure 4, the deviations of the experimental data measured in this work from the PR-CO₂ EoS using pure CO₂ were compared to the deviations of the Al-Siyabi et al work at various isotherms. The densities measured by Al-Siyabi et al. are only in the dense liquid / supercritical phase while in this work, the densities were measured in both gas and dense phases. Also, the deviations of the measured density of multi component mixture, MIX 1, from the PR-CO₂ equation of state can be seen at two isotherms 283.15 K and 323.15 K in Figure 5. The comparisons demonstrate that the density measurements are in good agreement with experimental data in literature as well as with prediction by equation of state.

3.5 Density measurement uncertainties

The combined standard uncertainties [42] of density measurements for each measured quantity were calculated using the root sum of the squares of uncertainties as shown in the following equation.

$$u_c(\rho) = \sqrt{u_1(T)^2 + u_2(p)^2 + u_3(\tau)^2 + u_4(x)^2} \quad (4)$$

where, $u_1(T)$ is the estimated uncertainties due to the temperature, $u_2(p)$ the estimated uncertainties due to pressure, $u_3(\tau)$ the estimated uncertainties of oscillation period and $u_4(x)$ the estimated uncertainties of the composition.

The estimated uncertainties due to temperature variations, $u_1(T)$, was calculated from the equation below:

$$u_1(T) = \sqrt{\left(\frac{\partial \rho}{\partial T}\right)^2 \cdot u(T)^2} \quad (5)$$

In the above equation, $u(T)$ is the standard estimated uncertainty of temperature probe and is considered to be 0.02 K [43]. The density gradient due to the temperature variations, $\left(\frac{\partial \rho}{\partial T}\right)$, was calculated from the equation below.

$$\left(\frac{\partial \rho}{\partial T}\right) = \frac{1}{2u(T)} (\rho_{T+u(T)} - \rho_{T-u(T)}) \quad (6)$$

The upper and lower limits of densities due to temperature effect, $\rho_{T+u(T)}$ and $\rho_{T-u(T)}$, were estimated from REFPROP v8.0 [40]. A similar procedure was followed to estimate the uncertainties due to the pressure and period of oscillation. The standard uncertainty of high pressure Quartzdyne pressure transducer (model: QS 30K-B), $u(p)$, is determined 0.02 MPa [39] and standard uncertainty of oscillation period, $u(\tau)$, is 0.005 μ s [44].

$$u_2(p) = \sqrt{\left(\frac{\partial \rho}{\partial p}\right)^2 \cdot u(p)^2} \quad (7)$$

$$\left(\frac{\partial \rho}{\partial p}\right) = \frac{1}{2u(p)} (\rho_{T+u(p)} - \rho_{T-u(p)}) \quad (8)$$

$$u_3(\tau) = \sqrt{\left(\frac{\partial \rho}{\partial \tau}\right)^2 \cdot u(\tau)^2} \quad (9)$$

$$\left(\frac{\partial \rho}{\partial \tau}\right) = \frac{1}{2u(\tau)} (\rho_{T+u(\tau)} - \rho_{T-u(\tau)}) \quad (10)$$

$$u_4(x) = \sqrt{\left(\frac{\partial \rho}{\partial x}\right)^2 \cdot u(x)^2} \quad (11)$$

$$\left(\frac{\partial \rho}{\partial x}\right) = \frac{1}{2u(x)} (\rho_{T+u(x)} - \rho_{T-u(x)}) \quad (12)$$

Finally, the expanded uncertainty of each measured density, $U(\rho)$, were calculated by multiplying to coverage factor, k . In this work, the coverage factor $k = 2$ was used to give a level of confidence of 95% for uncertainty of measurements.

$$U(\rho) = k u_c(\rho) \quad (13)$$

4. Modeling

The measured densities were used to evaluate the density predictions using classical cubic equations of state. In this work, SRK [18] and PR [19] equations of state were used to predict the densities. The modified binary interaction parameters, k_{ij} , shown in Table 5 and Table 6 for

PR [19] and SRK [18] equations of state, were employed to improve the phase equilibrium predictions. To improve the density predictions, also the CO₂ volume correction term [20][21] and Peneloux volume translation parameter [22] were applied to the SRK [18] and PR [19] equations of state.

Compressibility factors of the investigated mixtures were obtained from the real gas law, i.e. $Z=p/\rho RT$. Other thermodynamic properties such as isobaric heat capacity, speed of sound and Joule-Thomson coefficient were calculated from the Soave modifications [45][46] on the Benedict–Webb–Rubin (BWR) equation of state [47] by simplifying the original BWR equation as below:

$$Z = \frac{p}{\rho RT} = 1 + B\rho + D\rho^4 + E\rho^2(1 + F\rho^2) \exp(-F\rho^2) \quad (14)$$

In this work, the parameters B , D , E and F were optimized using the measured density data. The derived thermodynamic properties then were calculated [48] and compared to the GERG EoS [23]. The calculation procedure from the thermodynamic equations was explained in detail in the previous publication [48].

5. Results and discussions

First, the densitometer was calibrated using pure CO₂ at lower pressure conditions, i.e. in the gas phase and at elevated pressures, i.e. in the liquid and supercritical phases for each isotherms. Then, densities of three multi-component mixtures, MIX 1 with 0.0436 mole fraction impurity, MIX 2 with 0.1017 mole fraction impurity and MIX 3 with 0.3001 mole fraction impurity were measured at pressures ranging from 1 to 126 MPa at six different isotherms, $T/K = 273.15, 283.15, 298.15, 323.15, 373.15$ and 423.15 in the gas, liquid and supercritical regions. Both experimental results and modeling predictions using PR and SRK EoSs, those with the CO₂ correction volume and Peneloux shift parameter, as well as the GERG EoS, are shown in Table 7 through Table 9 and Figure 6 through Figure 13. In each table, the measured densities as well as the estimated uncertainties of measurements are shown at the corresponding pressure, temperature and phase. Also, the calculated density using PR EoS, that with CO₂ correction volume, with Peneloux shift parameter and the GERG EoS are presented in the tables. Finally, corresponding relative deviations of the models from experimental density for each measurement are also shown in each table. In addition, the Absolute Average Deviations (AADs) for all data are listed in the tables.

5.1. Measurement uncertainties

Table 10 shows the average and maximum estimated expanded uncertainties of density measurements in this work. The expanded uncertainties were reported with a level of confidence of 95% by multiplying the calculated combined standard uncertainty, $u_c(\rho)$, by coverage factor of $k = 2$ [42]. The average expanded uncertainties, $U(\rho)$, in the gas phase is 1.9% while in the dense liquid / supercritical phases is 0.4%. The maximum expanded uncertainty in the gas phase were reported 4.5% for MIX 1 at very low pressure and temperature while that in the dense liquid and supercritical phases is 0.9% and 1.3%, respectively. Generally, the uncertainties of measurements are higher at very low pressures in the gas phase as well as at points closed to the two-phase region in either the gas or liquid phase.

5.2. Evaluations of models

Table 11 summarizes the AAD and Maximum Absolute Deviation (MAD) of the models with PR and SRK equations of state using CO₂ volume correction, original and Peneloux shift parameters, from experimental data for each material at different regions. As can be seen, SRK-CO₂ equation of state predicts slightly more accurately compared to the PR-CO₂ EoS. The overall AADs for SRK and PR using CO₂ volume correction are 1.9% and 2.2%, respectively. The prediction accuracy of SRK-CO₂ in the gas, liquid and supercritical phases are higher than the PR-CO₂. The AAD for SRK-CO₂ in those phases are 2.9%, 1.6% and 1.3%, respectively. While those for the PR-CO₂ are 3.1%, 1.7% and 1.7%, respectively. The predictions of SRK and PR with modified k_{ij} without density correction parameters are almost in the same accuracy with the overall AAD of 4.7% and 4.8%, respectively. In the gas phase, the accuracy of SRK with the AAD of 2.8% is higher than PR with the AAD of 3.4%, while in the dense liquid / supercritical phase, the predictions of PR (AADs of 5.5% / 5.4%) is relatively better than the SRK (AADs of 5.8% / 5.7%). Using the Peneloux shift parameters to predict the density of CO₂ systems can result in an overall AAD of 3.0% and 4.1% for PR-Pen and SRK-Pen, respectively. The AADs in the gas phase are 2.8% and 3.2%, in the liquid phase are 4.5% and 2.6% and in the supercritical phase are 5.0% and 3.1% for SRK-Pen and PR-Pen, respectively.

By comparing the AADs of classical cubic equations of state, it is clear that CO₂ volume correction with PR and SRK equations of state can predict well compared to those without density correction term or those using Peneloux shift parameters for CO₂-rich mixtures. In the gas phase, SRK and SRK-Pen are the best models to predict the density with the AAD of 2.8%,

while in the dense liquid / supercritical phases, the SRK-CO₂ is the most accurate with the AADs of 1.6% / 1.3%. Overall, the SRK-CO₂ is the most accurate model to predict the density with the AAD of 1.9%.

The GERG multi parameter EoS was also evaluated by comparing predictions to the new experimental data measured in this work. As shown on [Table 11](#), the AAD of the GERG EoS in the gas, liquid and supercritical phases are 2.8%, 1.0% and 1.2%, respectively. The overall AAD of GERG EoS is also 1.7%. This means GERG EoS predicts the densities of investigated mixtures more accurately than the classical cubic EoSs which SRK-CO₂ with the AAD of 1.9% was the most accurate EoS.

5.3. Effect of impurities on the density reduction of pure CO₂

Increasing the density of CO₂ fluids will reduce the pipeline size and the running cost. However, the presence of common impurities in CO₂ stream will reduce the density of pure CO₂. The amount of reduction is a function of the mixture composition and pipeline operating pressure and temperature. The lighter components tend to reduce the density more. Also, the amount of reduction could be high at pressures and temperatures close to the critical pressure and temperature of the mixture. In order to investigate this effect, spline interpolation is implemented to the modeling and experimental data. [Table 12](#) and [Figure 12](#) to [Figure 14](#) show the reduction in CO₂ density for tests conducted at 323 K (50 °C). A maximum reduction of the CO₂ density at a certain pressure for a given temperature is observed for the CO₂ mixtures. The maximum reduction is 21.7% in MIX 1 ($MW = 43.64$), 33.9% in MIX 2 ($MW = 42.75$) and 38.5% in MIX 3 ($MW = 37.60$). The maximum reduction occurs at pressure around 11 MPa for MIX 1 and MIX 2 and at 12.4 MPa for MIX 3.

5.4. Derived thermodynamic properties

Thermodynamic properties were obtained from the measured density data. [Table 13](#) summarizes the AADs of the calculated properties from the GERG EoS [23] for all mixtures. [Figure 15](#) and [Figure 16](#) show the compressibility factor of MIX 1 versus density at various measured isotherms at total pressure ranges and lower pressure ranges, respectively. The obtained data then were compared to the prediction results using GERG EoS. [Figure 17](#) shows the relative deviations the GERG predictions from the calculated compressibility factor at different isotherms, while [Figure 18](#) shows the deviations for all three investigated mixtures. As can be seen, the predictions are in good agreement with the calculated data. The AADs of

GERG predictions from calculated compressibility factor for the investigated mixtures are 2.9% and 1.0% / 1.3% for the gas and dense liquid / supercritical phases with the overall AAD of 1.7%.

The derived specific heat capacities from the measured density data at various isotherms for the MIX 1 are shown in [Figure 19](#). The lines in the figure show the predictions using GERG EoS which is in good agreement with the experimentally calculated specific heat capacities. [Figure 20](#) shows the deviations of investigated mixtures from the GERG EoS. The overall AADs for MIX 1 to MIX 3 are 0.8%, 7.5% and 3.5%, respectively. The AAD for the investigated mixtures in the gas phase is 2.1%, while that of dense liquid / supercritical phases is 3.4% / 5.0%, respectively. The overall AAD of GERG predictions for the specific heat capacity of investigated mixtures is 3.5%.

[Figure 21](#) illustrates the calculated speed of sound as well as the predictions using GERG EoS for the mixture 1 at various isotherms. As can be seen, GERG EoS can predict the speed of sound accurately. The AADs for the mixture 1 in the gas, liquid and supercritical phases are 0.9%, 4.4% and 1.0% with the overall AAD of 2.1%, respectively. [Figure 22](#) also demonstrates the deviations of the calculated speed of sound for the three studied mixtures from the GERG EoS predictions. As can be seen in [Table 13](#), the AADs of speed of sound for the studied mixtures in the gas, liquid and supercritical phases are 0.5%, 6.8% and 2.6% with the overall AAD of 3.1%, respectively.

The Joule-Thomson coefficients, μ_{JT} , of MIX 1 at measured temperatures are shown in [Figure 23](#). The lines are predictions using GERG EoS, while the points show the calculated data from the measured densities. [Figure 24](#) shows the deviations of GERG predictions from the calculated data for all investigated mixtures. As can be seen in this figure, the μ_{JT} can reach the deviations about -50% for MIX 2 at 273 and 283 K in the area in the vicinity of bubble curve. The AAD of GERG from the calculated Joule-Thomson coefficient in the gas, liquid and supercritical phases are 5.2%, 10.3% and 10.0% with the overall AAD of 8.4%, respectively.

6. Conclusions

Undoubtedly, thermodynamic properties of CO₂ mixtures play an important role in the design and modeling of CCS infrastructures. This work was concentrated on the density measurement and modeling of various multi component mixtures. The densities of the mixtures were measured in the gas, liquid and supercritical regions after calibrations using pure CO₂ at each desired isotherm. It was concluded that the uncertainty of measurements in the gas phase is much higher than dense liquid / supercritical phases. The standard uncertainty of the pressure

transducer at lower pressures can result in higher uncertainty of the measured density in the gas phase. Also, it is obvious that the uncertainties of the measurement at few points which are in the vicinity of two-phase region are high due to the sharp changes of densities with pressure changes.

The importance of equations of state to predict the thermodynamic properties, particularly density, is evident. In the density modeling part of this work, two cubic equations of state, SRK and PR, as well as the GERG EoS were studied. These two classical cubic equations were selected due to the popularity in the oil and gas industry and availability in commercial software packages. Also, CO₂ volume correction was introduced to these cubic EoSs to improve the density prediction in the dense phase. It is concluded that generally both SRK and PR with CO₂ volume correction have acceptable predictions with AAD of 1.9% and 2.2%, respectively. The predictions by SRK in the gas phase are slightly more accurate than PR, while in the dense liquid / supercritical phase, predictions by PR are better than SRK. It also was concluded that introducing the CO₂ volume correction to the original equations improves the density predictions significantly in the dense liquid / supercritical phases. Overall, the AAD for PR was reduced from 4.8% to 2.2% by introducing the CO₂ volume correction. Those for SRK were reduced from 4.7% to 1.9%. The other conclusion worth noting is that the accuracy of predictions using CO₂ volume correction is higher than that using Peneloux shift parameter. It was also concluded that the GERG EoS is able to predict the densities of investigated mixture more accurately compared to the classical cubic EoSs. The overall AAD of GERG from the new measured data is 1.7%.

Moreover, the reduction in density of pure CO₂ due to the presence of impurities in the supercritical phase also was investigated for each mixture. It was concluded that a maximum reduction of the pure CO₂ density at a given temperature of 323.15 K (50 °C) was observed at pressures approximately 11 to 12 MPa depends on the composition of mixtures. Overall, lighter molecular weight impurities tend to reduce CO₂ density much more than those with a molecular weight close to pure CO₂.

Thermodynamic properties were also obtained from the measured density data. The calculated data then were used to evaluate the GERG EoS. The AADs of 1.7%, 3.5%, 3.8%, 3.1% and 8.4% were achieved for the Z-factor, isobaric and isochoric heat capacities, speed of sound and Joule-Thomson coefficient, respectively.

Acknowledgements

The JIP project "Impact of Common Impurities on Carbon Dioxide Capture, Transport and Storage" [20] was conducted jointly at Heriot-Watt University in UK and MINES ParisTech in France from 2011 - 2014. Thermophysical properties measured during the course of project were phase equilibria, hydrates [12], solid formation, density [21][49], viscosity, interfacial tension [50], solubility [51][52][53] and pH. This project was sponsored by Chevron, GALP Energia, Linde AG Engineering Division, OMV, Petroleum Expert, Statoil, TOTAL and National Grid Carbon Ltd, which is gratefully acknowledged.

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Table 1. Compositions of the multi-component mixtures along with the corresponding expanded uncertainties (coverage factor $k=2$) in the parentheses

Components	MIX 1 / mol%	MIX 2 / mol%	MIX 3 / mol%
Carbon Dioxide	95.6437	89.83	69.99
Methane	0.6261 (0.031)	0	20.02 (0.11)
Ethane	0	0	6.612 (0.034)
Propane	0	0	2.58 (0.013)
n-Butane	0	0	0.3997 (40 ppm)
i-Butane	0	0	0.3998 (40 ppm)
Nitrogen	1.41 (0.071)	5.05 (0.04)	0
Hydrogen	0.8175 (0.041)	0	0
Oxygen	0.08 (0.004)	3.07 (0.10)	0
Argon	1.21 (0.061)	2.05 (0.06)	0
Carbon Monoxide	0.2127 (0.011)	0	0
Total	100.00	100.00	100.00
Certification	BOC	BOC	BOC
Analysis Method	SM ^a	SM ^a	SM ^a

^a Supplier Method: The analysis techniques was based on the gas chromatography using flame ionization detector and thermal conductivity detector.

Table 2. Calibration data using pure CO₂ at low pressures (gas phase)

No	T/K	p/MPa	$\tau/\mu\text{s}$	$\rho_{REF}/\text{g}\cdot\text{cm}^{-3}$ [40]	$\tau^2/\mu\text{s}^2$
1	273.3	1.28	2417.372	0.0272	5843687.4
2	273.3	2.08	2419.303	0.0477	5853027.0
3	273.3	3.29	2423.028	0.0887	5871064.7
4	283.3	0.97	2420.026	0.0192	5856525.8
5	283.3	2.08	2422.275	0.0451	5867416.2
6	283.3	3.81	2427.179	0.0999	5891197.9
7	283.3	0.74	2419.422	0.0146	5853602.8
8	283.3	0.99	2419.942	0.0196	5856119.3
9	283.3	2.03	2422.182	0.0438	5866965.6
10	283.3	4.11	2428.783	0.1133	5898986.9
11	298.4	0.67	2423.921	0.0123	5875393.0
12	298.4	1.03	2424.438	0.0194	5877899.6
13	298.4	2.08	2426.469	0.0415	5887751.8
14	298.4	5.17	2435.391	0.1387	5931129.3
15	323.5	1.04	2432.229	0.0177	5915737.9
16	323.5	1.70	2433.385	0.0299	5921362.6
17	323.5	2.10	2434.084	0.0374	5924764.9
18	323.5	5.23	2440.921	0.1112	5958095.3
19	323.5	6.92	2446.297	0.1682	5984369.0
20	373.6	1.05	2448.685	0.0153	5996058.2
21	373.6	1.32	2449.071	0.0193	5997948.8
22	373.6	2.08	2450.061	0.0310	6002798.9
23	373.6	5.20	2455.071	0.0842	6027373.6
24	373.6	10.49	2466.093	0.2006	6081614.7
25	423.5	1.31	2465.475	0.0167	6078567.0
26	423.5	2.08	2466.363	0.0268	6082946.4
27	423.5	4.98	2470.217	0.0671	6101972.0
28	423.5	10.38	2478.314	0.1517	6142040.3
29	423.5	20.82	2497.709	0.3415	6238550.2

The standard uncertainty for the temperature, pressure and period are:
 $u(T) = 0.02$ K, $u(p) = 0.02$ MPa and $u(\tau) = 0.005$ μs

Table 3. Calibration data using pure CO₂ at high pressures (dense phase)

No	T/K	p/MPa	$\tau/\mu\text{s}$	$\rho_{REF}/\text{g.cm}^{-3}$ [40]	$\tau^2/\mu\text{s}^2$
1	273.1	3.61	2500.031	0.9292	6250155.0
2	273.1	4.96	2500.974	0.9407	6254870.9
3	273.1	10.76	2504.517	0.9786	6272605.4
4	273.1	20.94	2508.752	1.0243	6293836.6
5	273.1	52.06	2516.628	1.1088	6333416.5
6	273.1	103.63	2524.503	1.1907	6373115.4
7	273.1	125.17	2527.011	1.2161	6385784.6
8	283.0	4.60	2497.442	0.8637	6237216.5
9	283.0	5.20	2497.982	0.8724	6239914.1
10	283.0	10.48	2502.835	0.9249	6264183.0
11	283.0	21.33	2508.552	0.9868	6292833.1
12	283.0	51.44	2517.232	1.0803	6336456.9
13	283.0	104.13	2525.886	1.1707	6380100.1
14	283.0	125.56	2528.511	1.1973	6393367.9
15	298.5	4.99	2434.706	0.1304	5927793.3
16	298.5	12.34	2500.926	0.8472	6254630.9
17	298.5	20.40	2507.222	0.9152	6286162.2
18	298.5	50.96	2518.404	1.0358	6342358.7
19	298.5	76.41	2523.778	1.0928	6369455.4
20	298.5	103.65	2528.125	1.1380	6391416.0
21	298.5	124.91	2530.945	1.1668	6405682.6
22	298.5	22.90	2508.612	0.9302	6293134.2
23	323.7	40.12	2516.382	0.9221	6332178.4
24	323.7	25.52	2508.396	0.8361	6292050.5
25	323.7	76.90	2526.968	1.0352	6385567.3
26	323.7	104.17	2531.985	1.0878	6410948.0
27	323.7	125.53	2535.163	1.1205	6427051.4
28	323.7	125.53	2535.163	1.1205	6427051.4
29	373.5	17.16	2489.108	0.3984	6195658.6
30	373.5	20.84	2493.806	0.5002	6219068.4
31	373.5	35.63	2514.401	0.7198	6322212.4
32	373.5	52.12	2524.656	0.8286	6373887.9
33	373.5	104.28	2540.168	0.9914	6452453.5
34	373.5	125.86	2544.158	1.0325	6472739.9
35	423.4	26.99	2506.068	0.4470	6280376.8
36	423.4	51.34	2529.613	0.6953	6398941.9
37	423.4	75.91	2541.127	0.8158	6457326.4
38	423.4	104.24	2549.567	0.9024	6500291.9
39	423.4	124.66	2554.176	0.9487	6523815.0

The standard uncertainty for the temperature, pressure and period are:
 $u(T) = 0.02$ K, $u(p) = 0.02$ MPa and $u(\tau) = 0.005$ μs

Table 4. Calibration Parameters for Anton Paar DMA-HPM Densitometer using pure CO₂

<i>T</i> / K	Gas Phase		Liquid or Supercritical Phase	
	<i>A</i>	<i>B</i>	<i>A</i>	<i>B</i>
273.2	2.2489E-06	13.1148729	2.11433E-06	12.28403271
283.2	2.22815E-06	13.02872592	2.13096E-06	12.42475544
298.2	2.25362E-06	13.22771299	2.11904E-06	12.40548964
323.2	2.19724E-06	12.98059669	2.10897E-06	12.43289764
373.2	2.16346E-06	12.95646413	2.09829E-06	12.54758176
423.2	2.02721E-06	12.30358989	2.06409E-06	12.51471964

Table 5. Modified binary interaction parameters in this work for PR EoS [54]

	CO ₂	CO	N ₂	O ₂	Ar	H ₂	CH ₄	C ₂ H ₆	C ₃ H ₈	SO ₂	H ₂ S
CO ₂		-0.079	-0.014	0.111	0.129	0.089	0.099	0.129	0.131	0.02	0.082
CO			0.005	0 ^a	0.007	0 ^a	0.022	-0.003	0 ^a	0.024	0.085
N ₂				-0.013	-0.007	0 ^a	0.032	0.039	0.083	0.128	0.174
O ₂					0 ^a	0 ^a	0 ^a	0 ^a	0.112	0.222	0 ^a
Ar						0 ^a	0.026	0.054	0 ^a	0 ^a	0 ^a
H ₂							0 ^a	0 ^a	0 ^a	0 ^a	0 ^a
CH ₄								0.001	0.016	0.129	0.084
C ₂ H ₆									-0.006	0.11	0.084
C ₃ H ₈										0 ^a	0.082
SO ₂											0 ^a
H ₂ S											

^a The EoS was not tuned for this binary system

Table 6. Modified binary interaction parameters in this work for SRK EoS [54]

	CO ₂	CO	N ₂	O ₂	Ar	H ₂	CH ₄	C ₂ H ₆	C ₃ H ₈	SO ₂	H ₂ S
CO ₂		-0.062	-0.046	0.106	0.123	0.2	0.100	0.137	0.139	0.020	0.096
CO			0.006	0 ^a	0.008	0 ^a	0.030	-0.022	0 ^a	0.000	0.061
N ₂				-0.014	-0.008	0 ^a	0.030	0.032	0.078	0.091	0.157
O ₂					0 ^a	0 ^a	0 ^a	0 ^a	0.113	0.219	0 ^a
Ar						0 ^a	0.028	0.053	0 ^a	0 ^a	0 ^a
H ₂							0 ^a	0 ^a	0 ^a	0 ^a	0 ^a
CH ₄								-0.003	0.010	0.119	0.077
C ₂ H ₆									-0.005	0.11	0.087
C ₃ H ₈										0 ^a	0.087
SO ₂											0 ^a
H ₂ S											

^a The EoS was not tuned for this binary system

Table 7. Experimental and modeling results of MIX 1

Phase	T/K	p/MPa	$\rho_{exp}/kg.m^{-3}$	$u_c(\rho)/kg.m^{-3}$	$\rho_{model}/kg.m^{-3}$				$(\rho_{exp} - \rho_{model})/\rho_{exp} \times 100$			
					PR-CO ₂	PR	PR-Pen	GERG	PR-CO ₂	PR	PR-Pen	GERG
Gas	273.4	1.71	38.4	0.5	37.7	37.6	37.6	36.7	2.0	2.1	2.2	4.4
Gas	273.4	2.07	47.4	0.5	47.1	47.0	46.9	45.8	0.6	0.7	0.9	3.3
Gas	273.4	2.73	66.4	0.6	66.6	66.5	66.3	64.5	0.3	0.2	0.1	2.8
Liq.	273.4	6.71	888.0	3.3	894.0	893.8	859.8	881.8	0.7	0.6	3.2	0.7
Liq.	273.4	11.31	927.8	2.7	937.2	950.6	912.3	920.8	1.0	2.5	1.7	0.8
Liq.	273.4	21.80	983.9	2.2	997.9	1032.0	987.0	977.2	1.4	4.9	0.3	0.7
Liq.	273.4	36.27	1034.2	2.0	1051.3	1102.1	1050.8	1027.7	1.7	6.6	1.6	0.6
Liq.	273.4	51.73	1073.6	1.8	1092.6	1153.9	1097.9	1067.1	1.8	7.5	2.3	0.6
Liq.	273.4	76.40	1121.7	1.7	1142.2	1212.6	1150.9	1114.7	1.8	8.1	2.6	0.6
Liq.	273.4	104.38	1164.4	1.6	1185.1	1260.3	1193.8	1156.5	1.8	8.2	2.5	0.7
Liq.	273.4	126.02	1192.1	1.6	1211.9	1288.9	1219.4	1183.4	1.7	8.1	2.3	0.7
Gas	283.3	1.81	37.6	0.5	38.0	38.0	38.0	37.1	1.1	1.1	0.9	1.4
Gas	283.3	3.37	80.4	0.7	82.1	81.6	81.3	78.8	2.1	1.5	1.1	2.0
Liq.	283.3	6.36	810.8	0.7	797.9	778.3	752.4	821.4	1.6	4.0	7.2	1.3
Liq.	283.3	11.68	888.9	3.0	881.9	882.9	849.7	867.3	0.8	0.7	4.4	2.4
Liq.	283.3	22.57	956.3	2.3	960.9	987.7	946.4	941.0	0.5	3.3	1.0	1.6
Liq.	283.3	36.41	1004.6	2.0	1019.4	1065.2	1017.3	996.7	1.5	6.0	1.3	0.8
Liq.	283.3	54.13	1052.1	1.8	1070.7	1130.6	1076.8	1045.9	1.8	7.5	2.3	0.6
Liq.	283.3	78.00	1099.3	1.7	1121.2	1191.4	1131.8	1094.5	2.0	8.4	3.0	0.4
Liq.	283.3	105.09	1141.3	1.6	1164.8	1240.9	1176.3	1137.1	2.1	8.7	3.1	0.4
Liq.	283.3	124.85	1167.1	1.5	1190.8	1269.0	1201.6	1162.9	2.0	8.7	3.0	0.4
Gas	298.3	1.68	32.1	0.4	32.3	32.6	32.5	31.8	0.6	1.4	1.3	1.0
Gas	298.3	1.96	38.6	0.4	38.3	38.7	38.6	37.7	0.7	0.3	0.1	2.3
Gas	298.3	2.76	57.0	0.5	56.7	57.5	57.4	55.8	0.5	0.9	0.6	2.2
Gas	298.3	3.08	64.8	0.5	64.6	65.6	65.4	63.5	0.3	1.3	1.0	2.0
Liq.	298.3	12.55	778.4	3.6	784.9	767.0	741.9	772.8	0.8	1.5	4.7	0.7
Liq.	298.4	20.26	865.2	2.5	876.0	883.0	849.8	859.0	1.2	2.1	1.8	0.7
Liq.	298.4	50.12	999.6	1.8	1016.8	1068.2	1020.0	993.8	1.7	6.9	2.0	0.6
Liq.	298.4	75.62	1061.8	1.6	1080.1	1147.4	1092.0	1055.1	1.7	8.1	2.8	0.6
Liq.	298.4	103.39	1111.8	1.6	1130.0	1206.0	1144.9	1103.7	1.6	8.5	3.0	0.7
Liq.	298.4	126.33	1145.6	1.5	1162.8	1242.5	1177.8	1136.1	1.5	8.5	2.8	0.8
Gas	323.4	1.45	24.0	0.4	24.9	25.1	25.1	24.6	4.0	4.7	4.6	2.4
Gas	323.4	2.19	37.4	0.4	38.8	39.1	39.1	38.2	3.7	4.7	4.5	2.1
Gas	323.3	3.60	66.0	0.5	68.0	69.1	68.9	66.9	3.1	4.7	4.4	1.4
Gas	323.3	5.22	104.6	0.6	107.4	109.9	109.3	105.5	2.8	5.1	4.6	0.9
Gas	323.4	8.27	209.5	1.3	213.8	219.1	217.0	208.3	2.0	4.5	3.5	0.6
SC	323.4	15.88	642.4	3.8	649.7	617.0	600.7	633.5	1.1	3.9	6.5	1.4
SC	323.4	22.59	758.5	2.6	767.8	756.9	732.3	752.9	1.2	0.2	3.4	0.7
SC	323.4	29.52	822.1	2.2	833.7	840.6	810.5	817.2	1.4	2.3	1.4	0.6
SC	323.4	54.07	941.5	1.7	957.5	1001.4	958.9	936.8	1.7	6.4	1.9	0.5
SC	323.4	77.97	1008.7	1.6	1026.1	1088.4	1038.4	1003.1	1.7	7.9	2.9	0.6
SC	323.4	106.00	1065.5	1.5	1082.7	1157.1	1100.8	1058.3	1.6	8.6	3.3	0.7
SC	323.4	126.46	1099.1	1.5	1115.2	1194.9	1135.0	1090.4	1.5	8.7	3.3	0.8
Gas	373.5	2.12	29.1	0.3	31.3	31.5	31.4	30.8	7.7	8.4	8.3	6.1
Gas	373.5	2.75	40.2	0.3	41.3	41.6	41.5	40.6	2.8	3.6	3.4	1.2
Gas	373.5	3.52	50.3	0.3	53.7	54.2	54.1	52.8	6.6	7.7	7.4	5.0
Gas	373.5	5.24	79.9	0.4	83.6	84.7	84.4	82.2	4.6	6.0	5.6	2.9
Gas	373.5	10.61	191.4	0.7	196.2	198.3	196.6	192.4	2.5	3.6	2.7	0.5
SC	373.5	26.16	568.5	2.0	576.2	554.8	541.5	561.1	1.3	2.4	4.7	1.3
SC	373.5	53.57	805.3	1.6	815.5	836.6	806.8	799.8	1.3	3.9	0.2	0.7
SC	373.5	77.94	900.8	1.5	912.6	959.0	920.0	893.7	1.3	6.5	2.1	0.8
SC	373.6	104.58	970.6	1.4	982.9	1046.3	1000.1	961.4	1.3	7.8	3.0	1.0
SC	373.6	123.18	1008.7	1.4	1020.6	1092.1	1041.7	997.9	1.2	8.3	3.3	1.1
Gas	423.4	0.95	11.4	0.3	11.9	12.0	12.0	11.8	4.6	4.7	4.7	3.0

Gas	423.4	2.09	25.3	0.3	26.7	26.8	26.7	26.2	5.2	5.6	5.4	3.6
Gas	423.4	3.59	44.5	0.3	46.9	47.2	47.1	46.2	5.4	5.9	5.7	3.7
Gas	423.4	5.28	66.8	0.3	70.6	71.0	70.8	69.4	5.6	6.3	5.9	3.9
Gas	423.4	7.74	101.7	0.3	107.2	107.9	107.4	105.3	5.4	6.1	5.6	3.6
SC	423.4	34.18	521.3	1.3	527.1	514.5	503.0	515.3	1.1	1.3	3.5	1.1
SC	423.4	53.52	682.4	1.4	692.0	697.6	676.7	678.6	1.4	2.2	0.8	0.5
SC	423.4	76.14	791.7	1.3	803.6	832.9	803.3	787.7	1.5	5.2	1.5	0.5
SC	423.4	104.16	880.1	1.3	894.1	944.2	906.3	874.6	1.6	7.3	3.0	0.6
SC	423.4	121.85	922.5	1.2	937.2	996.5	954.5	915.7	1.6	8.0	3.5	0.7
<i>AAD / %</i>									2.1	4.9	3.0	1.5

The standard uncertainty for the temperature, pressure and period are:

$u(T) = 0.02$ K, $u(p) = 0.02$ MPa and $u(\tau) = 0.005$ μ s

Table 8. Experimental and modeling results of MIX 2

Phase	T/K	p/MPa	$\rho_{exp}/kg.m^{-3}$	$u_c(\rho)/kg.m^{-3}$	$\rho_{model}/kg.m^{-3}$				$(\rho_{exp} - \rho_{model})/\rho_{exp} \times 100$			
					PR-CO ₂	PR	PR-Pen	GERG	PR-CO ₂	PR	PR-Pen	GERG
Gas	273.2	1.79	38.4	0.5	38.4	38.3	38.2	37.9	0.2	0.3	0.7	1.4
Gas	273.2	2.24	49.8	0.5	50.0	50.0	49.7	49.3	0.5	0.3	0.1	1.0
Liq.	273.3	8.80	841.6	2.6	837.4	842.8	810.0	846.5	0.5	0.1	3.8	0.6
Liq.	273.3	10.92	867.7	2.2	865.7	876.1	840.8	871.0	0.2	1.0	3.1	0.4
Liq.	273.3	21.00	941.0	1.5	947.2	975.8	932.2	942.9	0.7	3.7	0.9	0.2
Liq.	273.3	52.04	1048.9	1.0	1064.0	1119.6	1062.7	1050.3	1.4	6.7	1.3	0.1
Liq.	273.3	104.16	1145.6	0.7	1163.2	1232.4	1163.9	1145.2	1.5	7.6	1.6	0.0
Liq.	273.3	125.71	1174.9	0.7	1191.4	1262.6	1190.8	1173.5	1.4	7.5	1.4	0.1
Gas	283.3	1.74	37.0	0.5	35.0	35.3	35.2	34.9	5.4	4.6	5.0	5.7
Gas	283.3	2.28	46.5	0.5	48.3	48.0	47.8	47.4	3.8	3.2	2.7	1.8
Liq.	283.3	10.67	803.7	2.9	783.5	781.6	753.2	793.6	2.5	2.7	6.3	1.3
Liq.	283.3	20.84	895.0	1.6	899.0	919.2	880.4	896.9	0.5	2.7	1.6	0.2
Liq.	283.3	52.46	1012.6	1.0	1036.0	1089.0	1035.1	1023.3	2.3	7.5	2.2	1.1
Liq.	283.3	104.12	1114.4	0.7	1141.3	1211.2	1144.9	1124.1	2.4	8.7	2.7	0.9
Liq.	283.3	125.26	1145.3	0.7	1170.9	1243.1	1173.4	1153.4	2.2	8.5	2.5	0.7
Gas	298.4	2.08	39.2	0.4	39.7	40.1	39.9	39.6	1.1	2.1	1.7	0.8
Gas	298.4	3.53	75.3	0.5	73.5	74.7	74.3	73.2	2.4	0.8	1.4	2.8
SC	298.4	12.58	702.1	3.5	684.0	670.9	649.5	693.6	2.6	4.4	7.5	1.2
SC	298.4	20.81	827.0	1.8	820.8	827.7	795.9	821.8	0.7	0.1	3.8	0.6
SC	298.4	51.34	979.9	1.0	988.3	1035.2	986.3	977.5	0.9	5.6	0.7	0.2
SC	298.4	104.10	1097.0	0.7	1108.9	1178.7	1115.9	1093.1	1.1	7.5	1.7	0.4
SC	298.4	125.88	1130.3	0.7	1141.8	1215.2	1148.5	1125.4	1.0	7.5	1.6	0.4
Gas	323.5	2.57	45.3	0.4	45.0	45.4	45.3	44.9	0.7	0.4	0.0	0.9
Gas	323.5	3.70	67.5	0.4	67.8	68.8	68.4	67.6	0.3	1.9	1.3	0.1
SC	323.5	12.27	409.1	2.7	406.6	383.8	375.8	387.6	0.6	6.2	8.1	5.3
SC	323.5	20.91	687.5	2.1	677.9	665.0	643.9	679.9	1.4	3.3	6.3	1.1
SC	323.5	51.56	909.3	1.0	914.3	950.9	909.3	906.7	0.5	4.6	0.0	0.3
SC	323.5	105.06	1047.7	0.7	1058.2	1126.1	1068.5	1044.5	1.0	7.5	2.0	0.3
SC	323.5	125.48	1082.4	0.7	1092.9	1166.0	1104.4	1078.4	1.0	7.7	2.0	0.4
Gas	373.6	1.53	20.7	0.3	21.7	21.8	21.8	21.7	4.7	5.2	4.9	4.6
Gas	373.6	2.57	34.5	0.3	37.3	37.6	37.4	37.2	8.1	8.8	8.4	7.9
SC	373.5	17.22	353.9	1.0	346.3	337.6	331.6	341.5	2.2	4.6	6.3	3.5
SC	373.5	21.00	439.7	1.2	435.9	419.7	410.7	429.6	0.9	4.6	6.6	2.3
SC	373.5	52.56	777.7	1.0	776.8	794.1	764.6	773.5	0.1	2.1	1.7	0.5
SC	373.6	103.94	951.8	0.7	958.3	1015.8	968.5	947.7	0.7	6.7	1.8	0.4
SC	373.6	125.15	996.1	0.6	1003.1	1069.3	1017.1	990.9	0.7	7.4	2.1	0.5
Gas	423.5	2.40	30.4	0.3	30.0	30.1	30.0	29.9	1.2	0.8	1.2	1.4
Gas	423.5	3.41	43.6	0.3	43.3	43.5	43.3	43.2	0.6	0.1	0.5	0.8
SC	423.4	18.47	283.7	0.5	273.7	270.5	266.5	272.3	3.5	4.7	6.1	4.0
SC	423.4	21.75	335.5	0.6	327.0	320.8	315.4	324.9	2.5	4.4	6.0	3.2
SC	423.4	52.79	660.6	0.8	657.9	662.1	641.3	655.5	0.4	0.2	2.9	0.8
SC	423.4	102.77	862.5	0.6	867.4	912.0	873.5	858.9	0.6	5.7	1.3	0.4
SC	423.4	124.65	915.9	0.6	922.2	977.5	933.5	911.1	0.7	6.7	1.9	0.5
<i>AAD / %</i>									1.6	4.3	2.9	1.4

The standard uncertainty for the temperature, pressure and period are:

$$u(T) = 0.02 \text{ K}, u(p) = 0.02 \text{ MPa} \text{ and } u(\tau) = 0.005 \mu\text{s}$$

Table 9. Experimental and modeling results of MIX 3

Phase	T/K	p/MPa	$\rho_{exp}/kg.m^{-3}$	$u_c(\rho)/kg.m^{-3}$	$\rho_{model}/kg.m^{-3}$				$(\rho_{exp}-\rho_{model})/\rho_{exp} \times 100$			
					PR-CO ₂	PR	PR-Pen	GERG	PR-CO ₂	PR	PR-Pen	GERG
Gas	273.3	1.07	20.0	0.4	19.0	19.0	19.0	18.9	4.8	4.8	5.0	5.5
Gas	273.3	2.13	41.6	0.5	41.3	41.3	41.1	40.8	0.6	0.7	1.1	1.9
Liq.	273.2	12.71	685.5	1.5	667.5	674.7	641.7	665.8	2.6	1.6	6.4	2.9
Liq.	273.2	20.89	740.2	1.4	738.5	753.7	712.7	723.2	0.2	1.8	3.7	2.3
Liq.	273.2	52.28	838.1	1.3	858.1	889.7	833.2	824.2	2.4	6.2	0.6	1.7
Liq.	273.2	103.76	920.2	1.3	947.6	987.6	918.5	906.3	3.0	7.3	0.2	1.5
Liq.	273.2	124.50	944.3	1.3	971.3	1,012.6	940.1	929.8	2.9	7.2	0.5	1.5
Gas	283.3	1.12	18.7	0.4	19.1	19.2	19.2	19.1	2.1	2.5	2.3	1.7
Gas	283.3	2.09	36.8	0.4	38.0	38.2	38.0	37.7	3.3	3.8	3.5	2.5
Gas	283.3	4.87	110.8	0.8	117.7	117.1	116.1	114.7	6.2	5.6	4.7	3.4
Liq.	283.3	9.52	564.7	2.1	531.5	529.3	508.8	544.2	5.9	6.3	9.9	3.6
Liq.	283.3	20.63	701.4	1.3	697.0	707.5	671.3	686.1	0.6	0.9	4.3	2.2
Liq.	283.3	51.89	814.3	1.3	834.0	864.0	810.6	802.2	2.4	6.1	0.5	1.5
Liq.	283.3	103.37	902.8	1.3	930.4	970.8	903.9	890.2	3.1	7.5	0.1	1.4
Liq.	283.3	125.37	929.6	1.3	957.1	999.1	928.5	916.3	3.0	7.5	0.1	1.4
Gas	298.4	1.10	18.6	0.3	17.6	17.7	17.7	17.6	5.4	5.0	5.1	5.8
Gas	298.4	2.09	36.0	0.4	35.2	35.4	35.3	35.0	2.3	1.6	1.8	2.8
Gas	298.4	5.18	108.2	0.6	108.0	109.9	108.9	107.2	0.2	1.5	0.7	0.9
SC	298.4	10.96	462.0	2.0	438.7	431.4	417.7	445.3	5.0	6.6	9.6	3.6
SC	298.3	20.83	642.9	1.3	636.7	640.5	610.7	631.1	1.0	0.4	5.0	1.8
SC	298.3	51.73	779.9	1.2	799.4	826.4	777.5	770.6	2.5	6.0	0.3	1.2
SC	298.3	102.11	873.8	1.2	904.2	944.7	881.2	865.8	3.5	8.1	0.9	0.9
SC	298.3	124.97	903.9	1.3	934.4	977.3	909.6	894.9	3.4	8.1	0.6	1.0
Gas	323.4	1.22	19.3	0.3	17.8	17.9	17.9	17.8	7.6	7.2	7.3	7.9
Gas	323.5	2.11	32.7	0.3	32.0	32.2	32.1	31.9	2.3	1.6	1.8	2.7
Gas	323.4	5.23	90.6	0.5	90.9	92.3	91.7	90.3	0.3	1.9	1.2	0.3
SC	323.5	11.81	316.8	1.1	308.5	296.1	289.6	297.6	2.6	6.5	8.6	6.1
SC	323.5	20.23	531.8	1.1	520.9	513.0	493.7	519.3	2.0	3.5	7.2	2.3
SC	323.4	49.82	718.0	1.1	735.2	755.2	714.1	712.1	2.4	5.2	0.5	0.8
SC	323.4	103.43	837.2	1.2	866.4	906.0	847.5	830.8	3.5	8.2	1.2	0.8
SC	323.5	124.84	868.2	1.2	898.0	940.9	878.0	860.7	3.4	8.4	1.1	0.9
Gas	373.5	2.10	24.5	0.3	26.6	26.8	26.7	26.5	8.8	9.3	9.0	8.4
Gas	373.6	5.24	68.9	0.3	71.3	72.0	71.6	70.9	3.5	4.5	3.9	2.9
Gas	373.6	10.43	156.3	0.4	159.4	160.7	158.7	158.4	1.9	2.8	1.5	1.3
SC	373.6	20.73	364.2	0.7	359.6	349.7	340.6	353.9	1.2	4.0	6.5	2.8
SC	373.6	46.52	601.3	0.9	608.4	614.1	586.6	593.9	1.2	2.1	2.4	1.2
SC	373.5	63.34	668.4	1.0	684.5	701.9	666.2	663.4	2.4	5.0	0.3	0.7
SC	373.5	63.16	667.7	1.0	683.8	701.1	665.5	662.8	2.4	5.0	0.3	0.7
SC	373.6	103.40	767.8	1.1	792.4	826.6	777.6	761.5	3.2	7.7	1.3	0.8
SC	373.6	122.28	800.7	1.1	826.6	865.5	812.0	793.3	3.2	8.1	1.4	0.9
SC	373.6	125.29	804.9	1.1	831.4	871.0	816.8	797.9	3.3	8.2	1.5	0.9
Gas	423.4	2.12	23.2	0.2	23.3	23.4	23.3	23.2	0.7	1.0	0.8	0.4
Gas	423.4	5.22	58.2	0.3	59.8	60.1	59.8	59.5	2.8	3.3	2.8	2.3
Gas	423.4	10.36	120.5	0.3	125.9	126.2	125.0	124.9	4.5	4.8	3.8	3.7
Gas	423.4	20.94	258.3	0.5	271.3	267.7	262.3	268.3	5.0	3.6	1.6	3.9
SC	423.5	29.08	371.6	0.6	370.1	363.0	353.3	364.4	0.4	2.3	4.9	2.0
SC	423.5	49.87	527.2	0.8	534.5	535.3	514.3	522.2	1.4	1.5	2.4	1.0
SC	423.5	103.36	704.5	0.9	725.8	753.1	712.3	698.8	3.0	6.9	1.1	0.8
SC	423.5	124.80	746.2	1.0	770.1	803.8	757.4	739.5	3.2	7.7	1.5	0.9
AAD / %									3.0	5.0	2.9	2.2

The standard uncertainty for the temperature, pressure and period are:

$$u(T) = 0.02 \text{ K}, u(p) = 0.02 \text{ MPa and } u(\tau) = 0.005 \text{ } \mu\text{s}$$

Table 10. Uncertainties of density measurements for each mixture with 95% level of confidence ($k=2$)

Material	Phase	No	Average		Max	
			$U(\rho)/\text{kg.m}^{-3}$	$U(\rho)/\%$	$U(\rho)/\text{kg.m}^{-3}$	$U(\rho)/\%$
MIX 1	Gas	24	0.94	1.8	2.64	4.5
	Liquid	22	4.01	0.4	7.28	0.9
	SC	17	3.43	0.4	7.58	1.2
	Total	63	2.68	1.0	7.58	4.5
MIX 2	Gas	12	0.81	2.0	1.06	2.9
	Liquid	11	2.83	0.3	5.80	0.7
	SC	20	2.28	0.4	7.07	1.3
	Total	43	2.01	0.8	7.07	2.9
MIX 3	Gas	18	0.79	1.9	1.62	3.9
	Liquid	10	2.79	0.4	4.10	0.7
	SC	21	2.16	0.3	3.92	0.8
	Total	49	1.74	0.9	4.10	3.9
Total	Gas	54	0.86	1.9	2.64	4.5
	Liquid	43	3.43	0.4	7.28	0.9
	SC	58	2.57	0.4	7.58	1.3
	Total	155	2.20	0.9	7.58	4.5

Table 11. AAD and Max. Deviations of this work using PR, SRK and GERG EoSs

Mixture	Phase	No	AAD / %							MAD / %						
			PR-CO ₂	SRK-CO ₂	PR	SRK	PR-Pen	SRK-Pen	GERG	PR-CO ₂	SRK-CO ₂	PR	SRK	PR-Pen	SRK-Pen	GERG
MIX 1	Gas	24	3.1	3.2	3.8	2.8	3.5	3.0	2.6	7.7	7.6	8.4	10.2	8.3	7.7	6.1
	Liquid	22	1.5	1.2	5.9	5.4	2.7	4.6	0.8	2.1	1.9	8.7	13.8	7.2	9.3	2.4
	SC	17	1.4	0.7	5.3	5.0	2.8	4.8	0.8	1.7	1.2	8.7	12.0	6.5	7.9	1.4
	Total	63	2.1	1.8	4.9	4.3	3.0	4.1	1.5	7.7	7.6	8.7	13.8	8.3	9.3	6.1
MIX 2	Gas	12	2.4	2.3	2.4	2.5	2.3	2.4	2.4	8.1	7.8	8.8	7.5	8.4	7.8	7.9
	Liquid	11	1.4	1.2	5.2	5.9	2.5	4.4	0.5	2.5	3.3	8.7	12.0	6.3	7.4	1.3
	SC	20	1.2	1.3	5.1	6.6	3.5	5.4	1.3	3.5	4.4	7.7	12.4	8.1	9.4	5.3
	Total	43	1.6	1.6	4.3	5.3	2.9	4.3	1.4	8.1	7.8	8.8	12.4	8.4	9.4	7.9
MIX 3	Gas	18	3.5	3.0	3.6	2.9	3.2	2.8	3.2	8.8	8.2	9.3	8.0	9.0	8.2	8.4
	Liquid	10	2.6	2.8	5.2	6.4	2.6	4.3	2.0	5.9	8.7	7.5	14.0	9.9	10.8	3.6
	SC	21	2.6	1.9	5.7	5.5	2.8	4.7	1.5	5.0	7.2	8.4	13.0	9.6	10.3	6.1
	Total	49	3.0	2.4	5.0	4.6	2.9	3.9	2.2	8.8	8.7	9.3	14.0	9.9	10.8	8.4
Total	Gas	54	3.1	2.9	3.4	2.8	3.2	2.8	2.8	8.8	8.2	9.3	10.2	9.0	8.2	8.4
	Liquid	43	1.7	1.6	5.5	5.8	2.6	4.5	1.0	5.9	8.7	8.7	14.0	9.9	10.8	3.6
	SC	58	1.7	1.3	5.4	5.7	3.1	5.0	1.2	5.0	7.2	8.7	13.0	9.6	10.3	6.1
	Total	155	2.2	1.9	4.8	4.7	3.0	4.1	1.7	8.8	8.7	9.3	14.0	9.9	10.8	8.4

Table 12. Density reduction of pure CO₂ at supercritical area at the temperature of 323.15 K

Material	<i>Min</i> / %	<i>p</i> / MPa
MIX 1	-20.1	11.01
MIX 2	-33.9	11.01
MIX 3	-45.9	12.39

Table 13. AADs of the calculated thermodynamic properties from the GERG EoS

Mixture	Phase	No	AAD / %				
			ρ	Z	C_p	SoS	μ_{JT}
MIX 1	Gas	24	3.0	2.9	1.2	0.9	4.0
	Liquid	22	0.8	0.8	0.6	4.4	3.1
	SC	17	0.8	0.8	0.7	1.0	4.8
	Total	63	1.6	1.6	0.8	2.1	3.9
MIX 2	Gas	12	2.5	2.4	4.1	0.3	12.3
	Liquid	11	0.5	0.5	7.5	15.1	26.5
	SC	20	1.3	1.4	9.6	5.6	11.5
	Total	43	1.4	1.4	7.5	6.6	15.6
MIX 3	Gas	18	3.2	3.2	1.9	0.2	2.0
	Liquid	10	2.0	2.1	5.2	3.1	8.3
	SC	21	1.5	1.6	4.0	1.1	12.7
	Total	49	2.3	2.2	3.5	1.1	7.9
Total	Gas	54	2.9	2.9	2.1	0.5	5.2
	Liquid	43	1.0	1.0	3.4	6.8	10.3
	SC	58	1.2	1.3	5.0	2.6	10.0
	Total	155	1.8	1.7	3.5	3.1	8.4

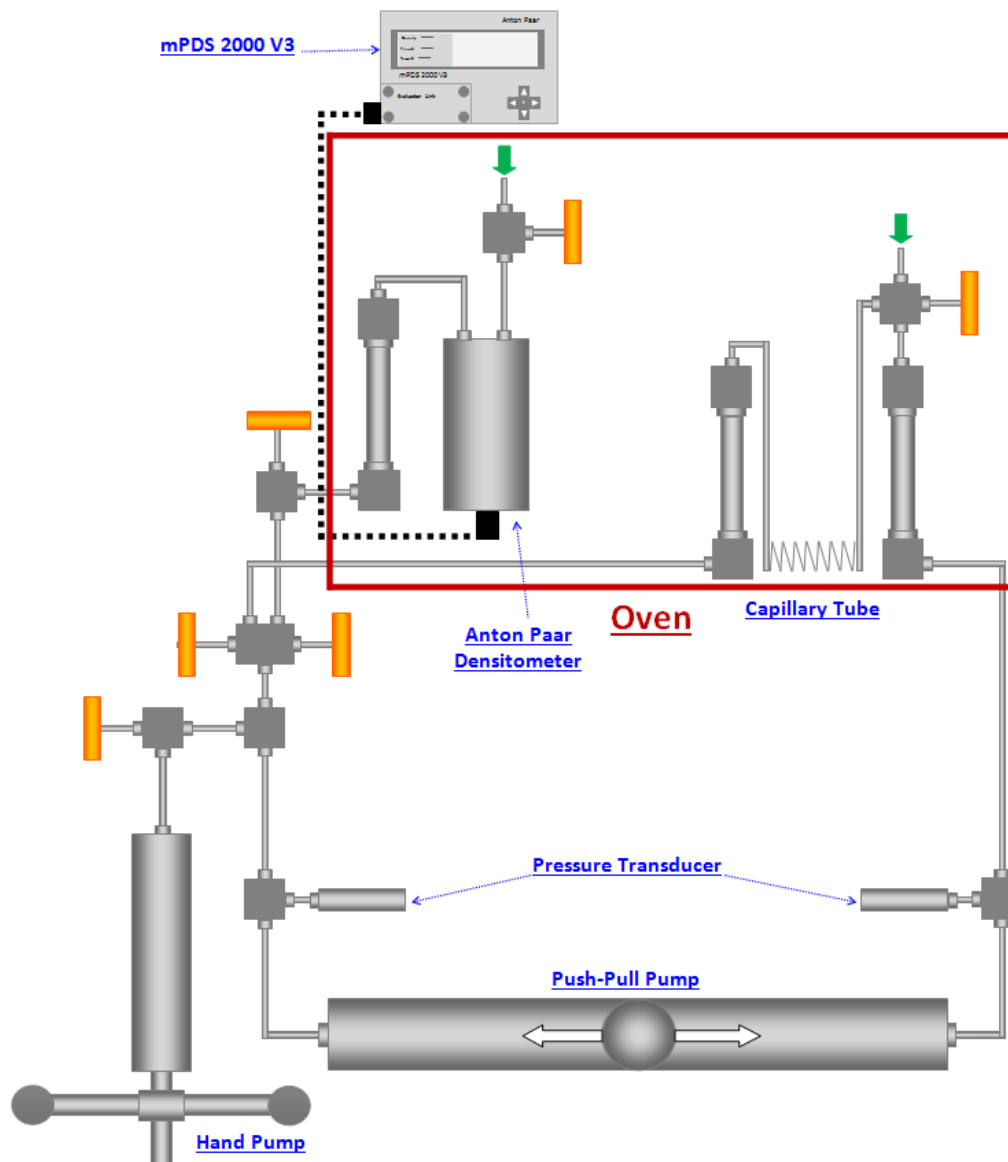


Figure 1. A schematic view of the viscosity experimental setup

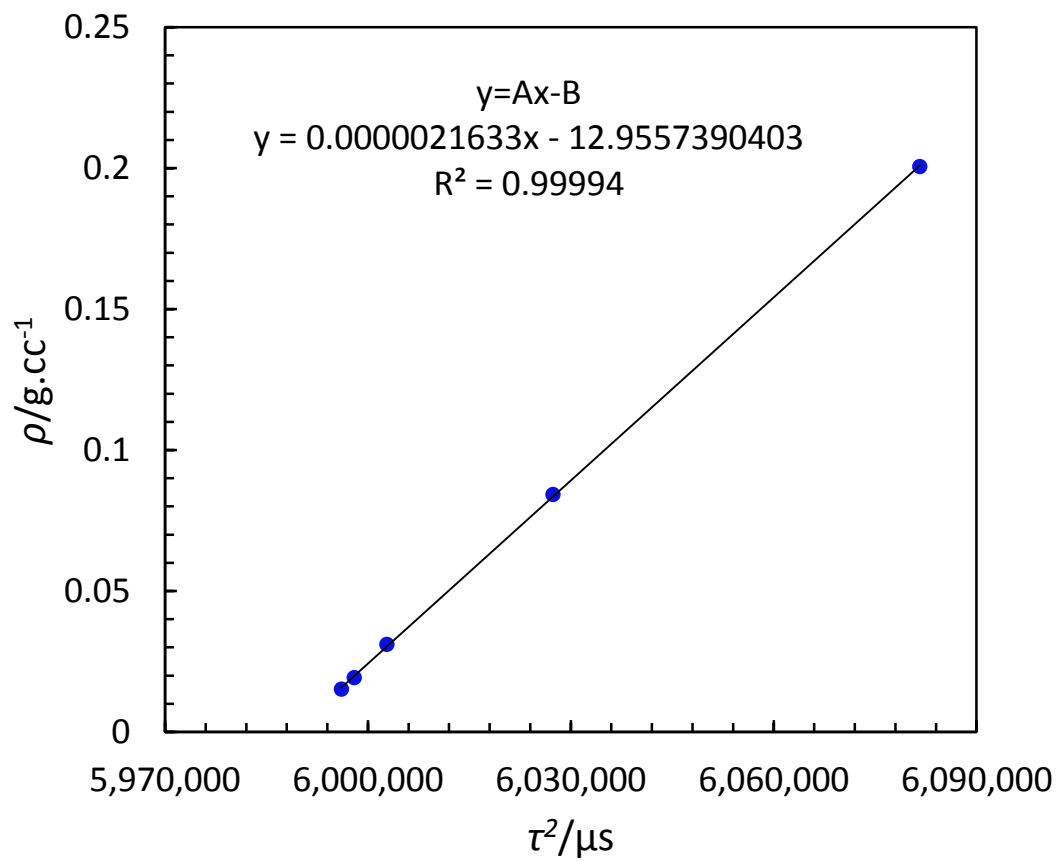


Figure 2. Calibration procedure using pure CO₂ at 373.15 K at low pressures (gas phase)

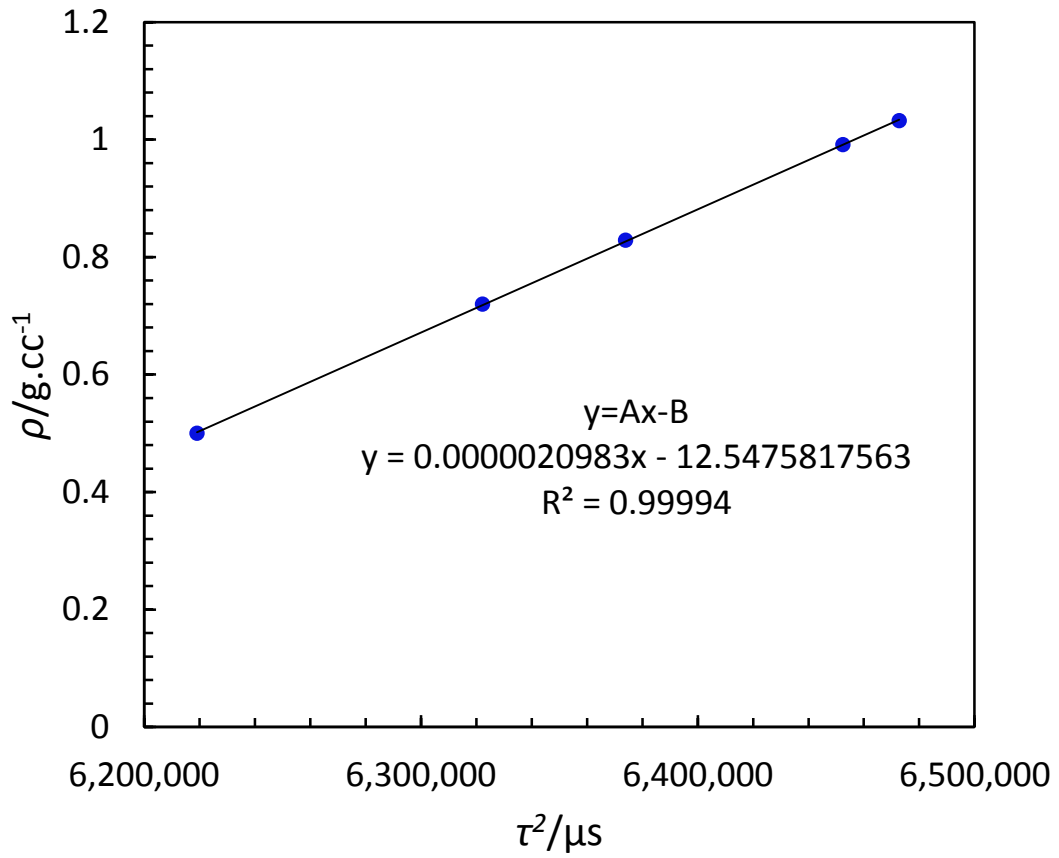


Figure 3. Calibration procedure using pure CO₂ at 373.15 K at high pressures (dense phase)

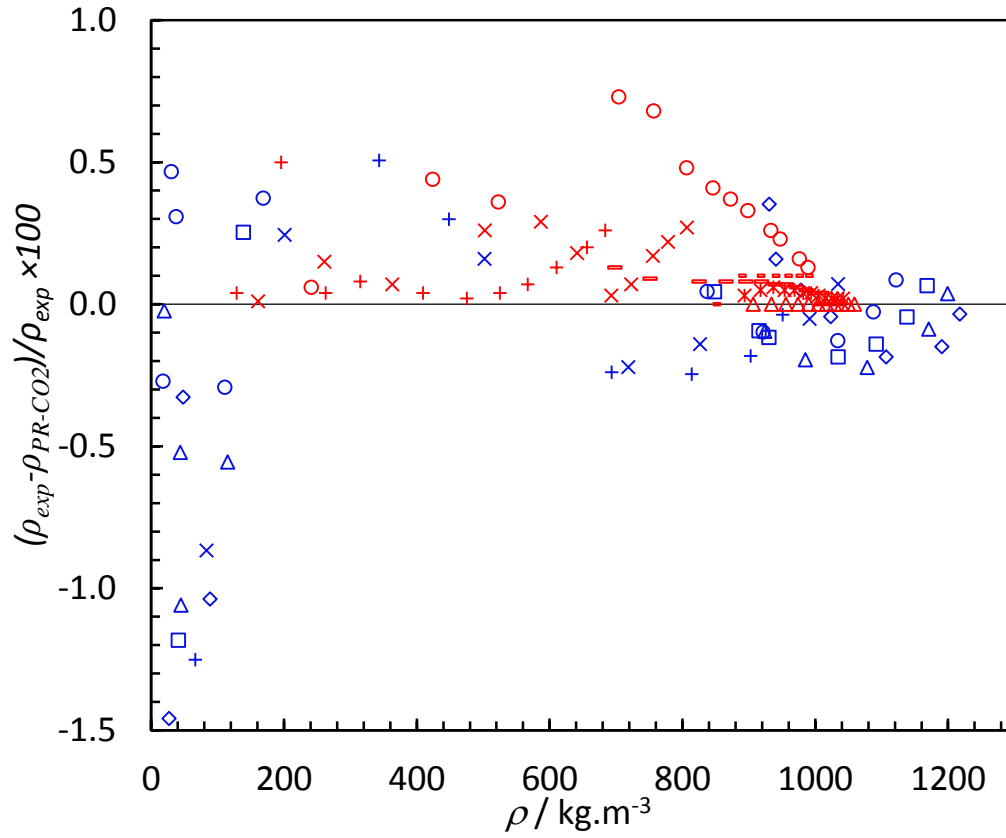


Figure 4. Validation data of pure CO₂ density at different isotherms

This work: (\diamond): 273.2 K, (Δ): 283.2 K, (\square): 298.2 K, (\circ): 323.2 K, (\times): 373.2 K, (+): 423.2 K
 Al-Siyabi et al. [41]: (Δ): 283.2 K, (\circ): 288.2 K, (-): 293.2 K, ($_$): 301.2 K, (\circ): 323.15 K, (\times):
 373.15 K, (+): 423.15 K

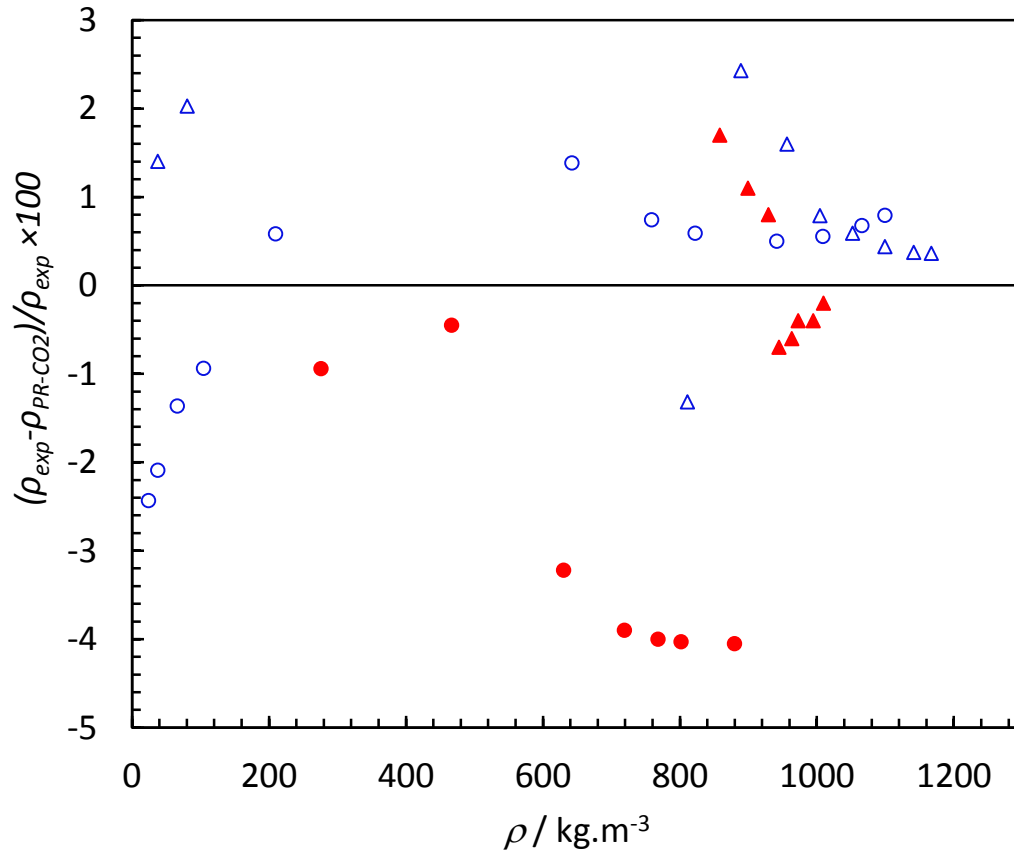


Figure 5. Density validation data at two isotherms for MIX 1

This work: (Δ): 283.2 K, (\circ): 323.2 K, Al-Siyabi et al. [41]: (\blacktriangle): 283.2 K, (\bullet): 323.2 K

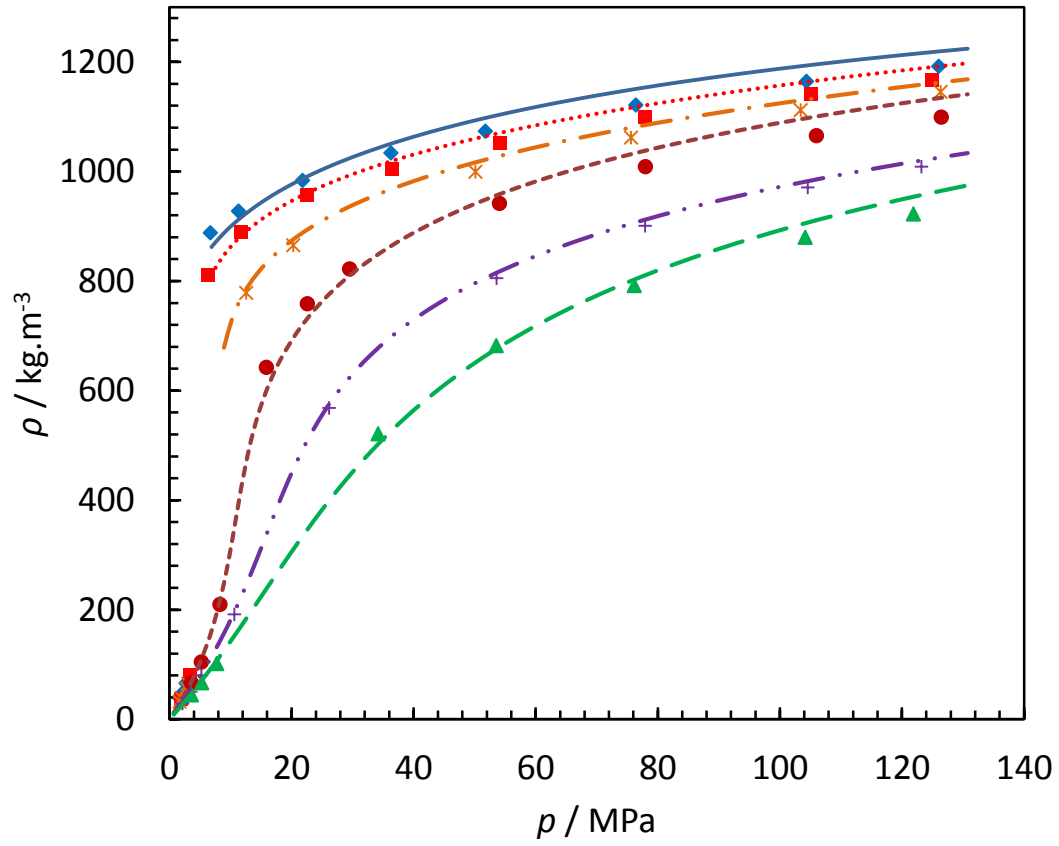


Figure 6. Experimental and modeling results of MIX 1
 Experimental / modeling (PR-CO₂) results: (♦/—) at 273.15 K, (■/.....) at 283.15 K, (x/—.)
 at 298.15 K, (●/----) at 323.15 K, (+/—...) at 373.15 K and (▲/—) at 423.15 K

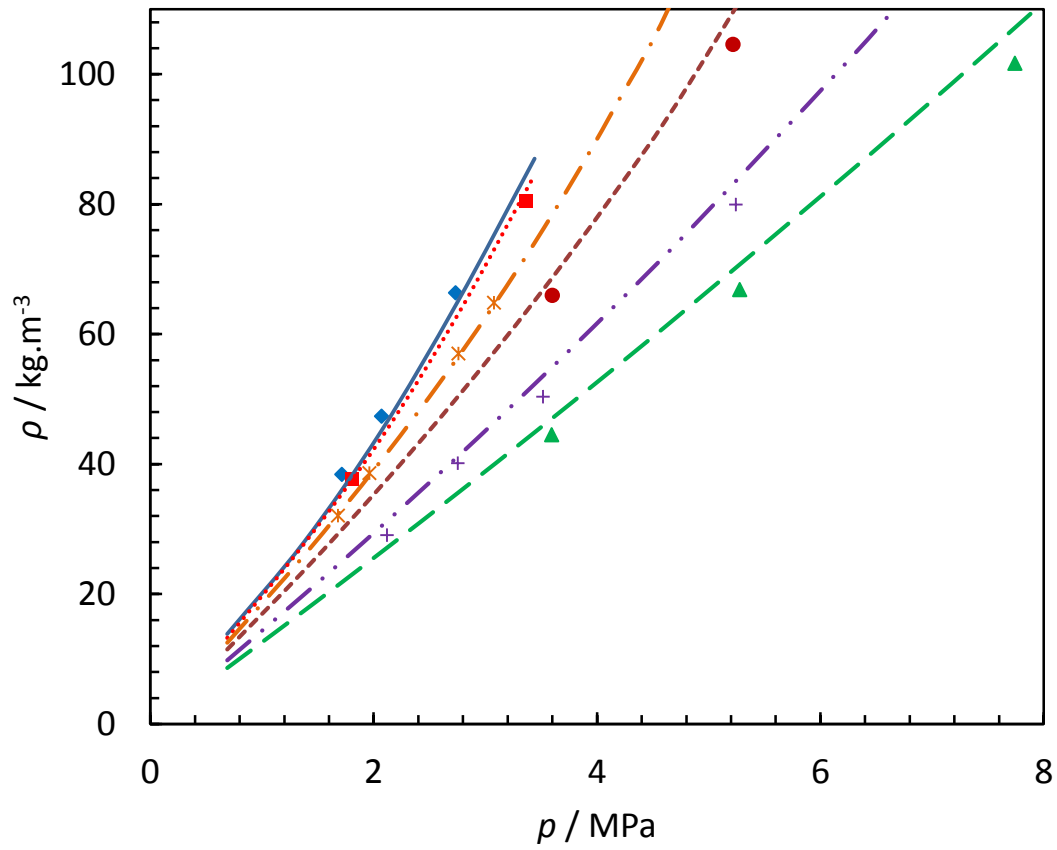


Figure 7. Experimental and modeling results of MIX 1 at lower pressures
 Experimental / modeling (PR-CO₂) results: (♦/—) at 273.15 K, (■/.....) at 283.15 K, (x/—.)
 at 298.15 K, (●/---) at 323.15 K, (+/—...) at 373.15 K and (▲/—) at 423.15 K

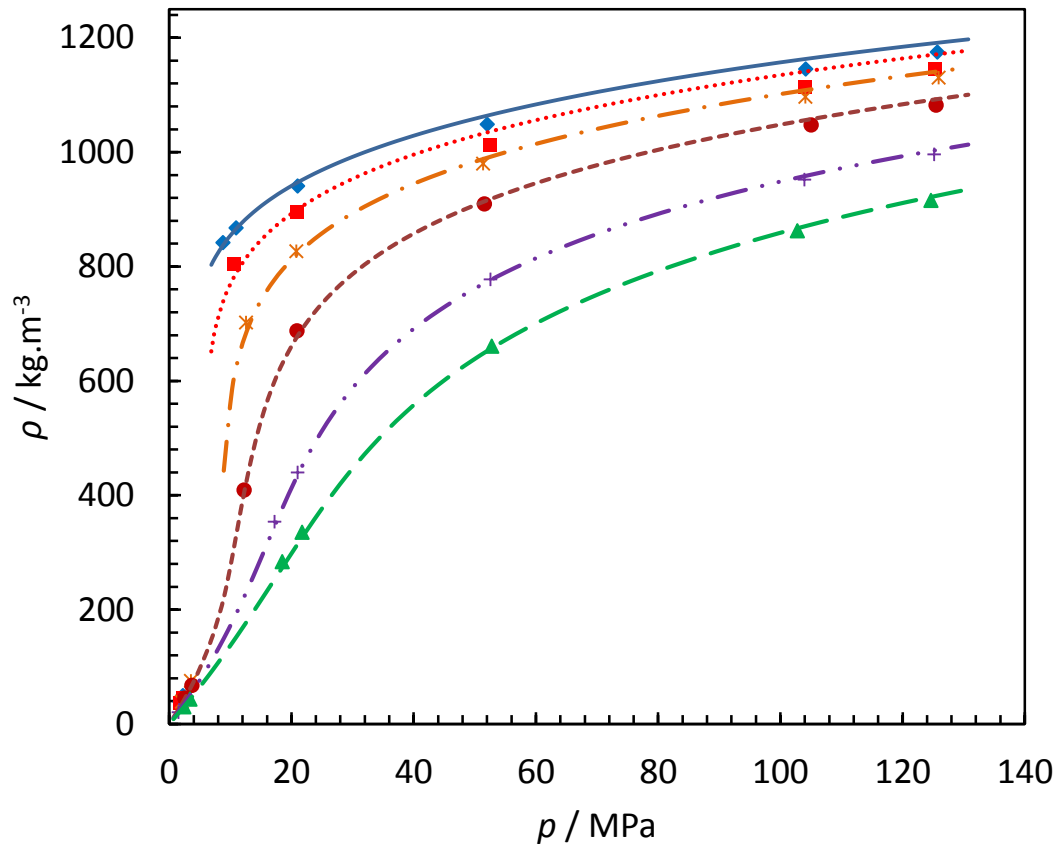


Figure 8. Experimental and modeling results of MIX 2
 Experimental / modeling (PR-CO₂) results: (♦/—) at 273.15 K, (■/.....) at 283.15 K, (x/—) at 298.15 K, (●/—) at 323.15 K, (+/—) at 373.15 K and (▲/—) at 423.15 K

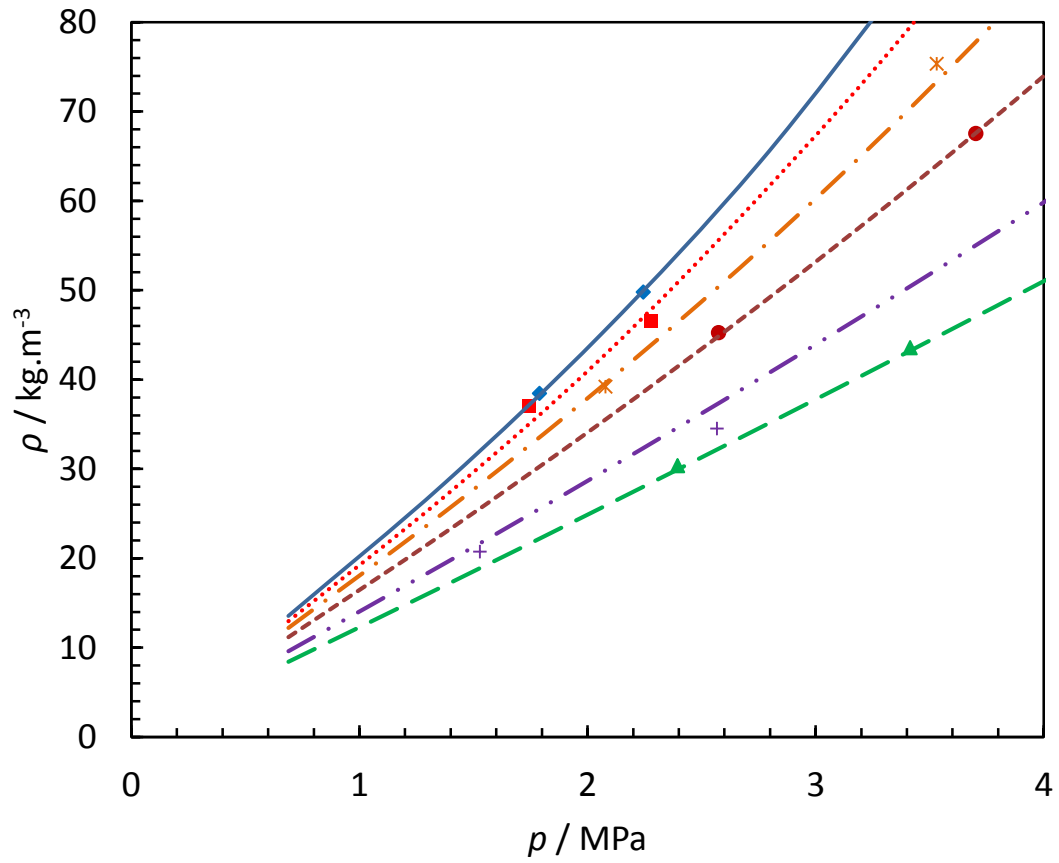


Figure 9. Experimental and modeling results of MIX 2 at low pressures
 Experimental / modeling (PR-CO₂) results: (◆/—) at 273.15 K, (■/⋯) at 283.15 K, (x/—·) at 298.15 K, (●/- - -) at 323.15 K, (+/—··) at 373.15 K and (▲/— - -) at 423.15 K

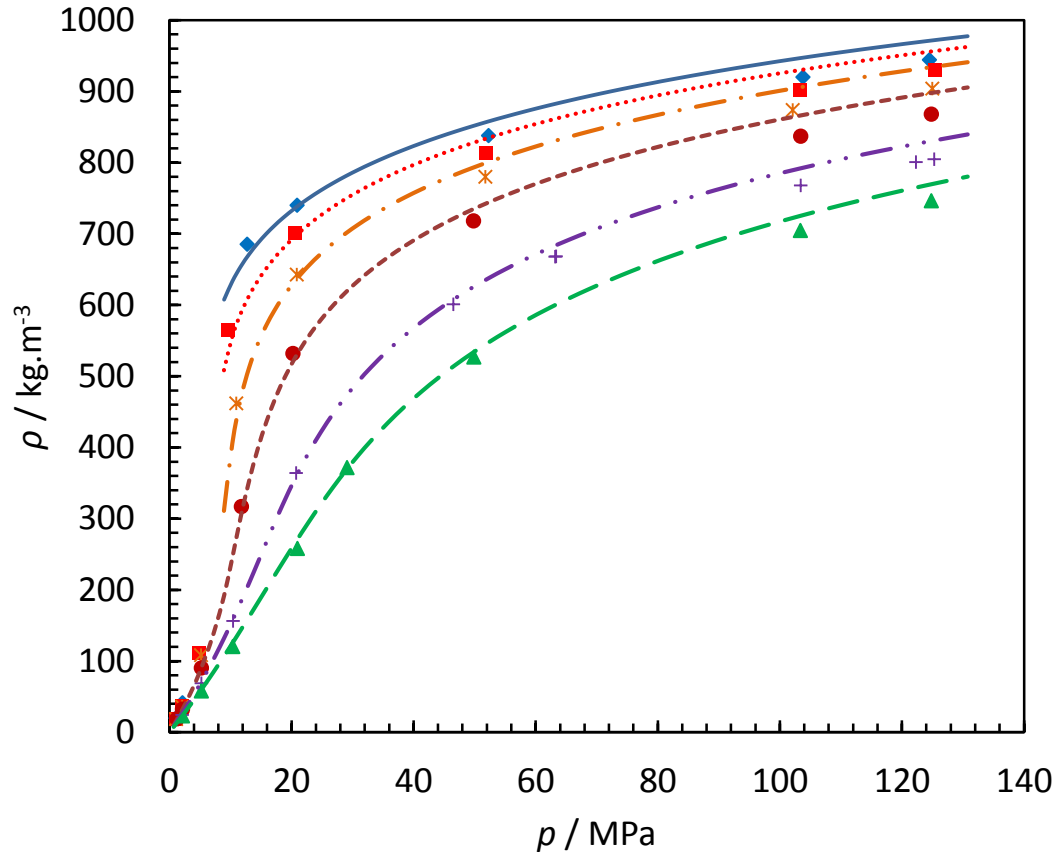


Figure 10. Experimental and modeling results of MIX 3
 Experimental / modeling (PR-CO₂) results: (♦/—) at 273.15 K, (■/.....) at 283.15 K, (x/—.) at 298.15 K, (●/---) at 323.15 K, (+/—...) at 373.15 K and (▲/—) at 423.15 K

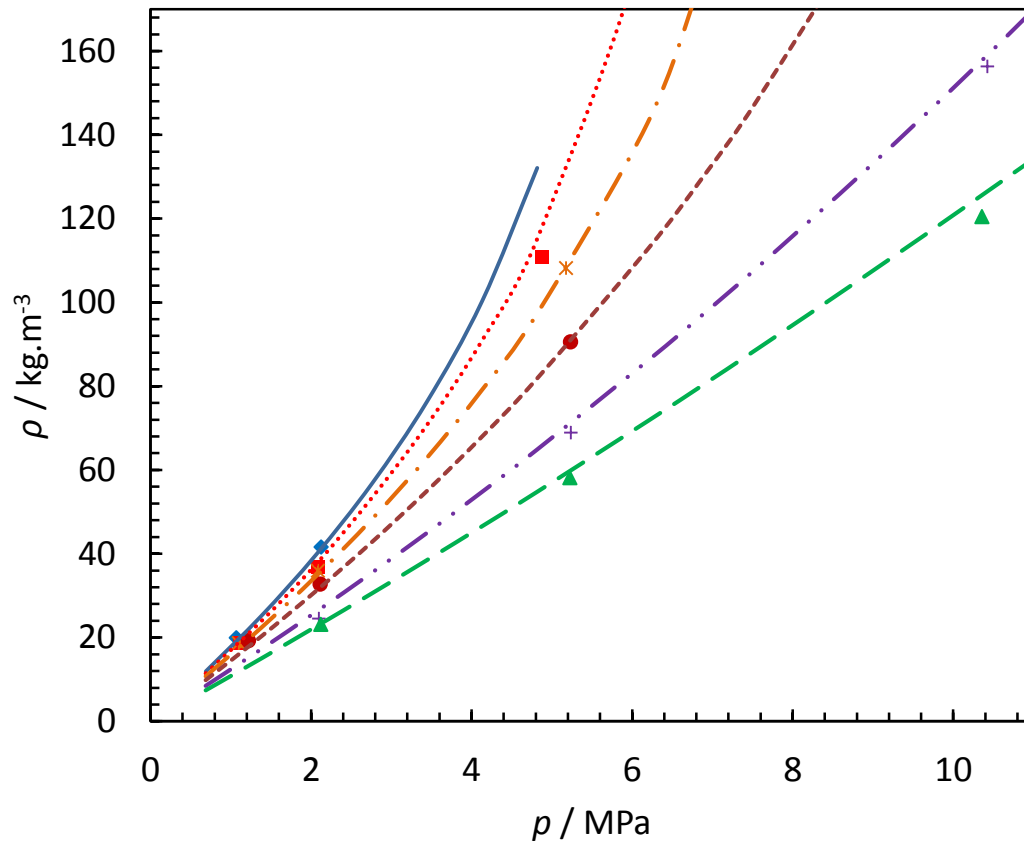


Figure 11. Experimental and modeling results of MIX 3 at low pressures
 Experimental / modeling (PR-CO₂) results: (◆/—) at 273.15 K, (■/.....) at 283.15 K, (x/—.)
 at 298.15 K, (●/- - -) at 323.15 K, (+/- ···) at 373.15 K and (▲/— —) at 423.15 K

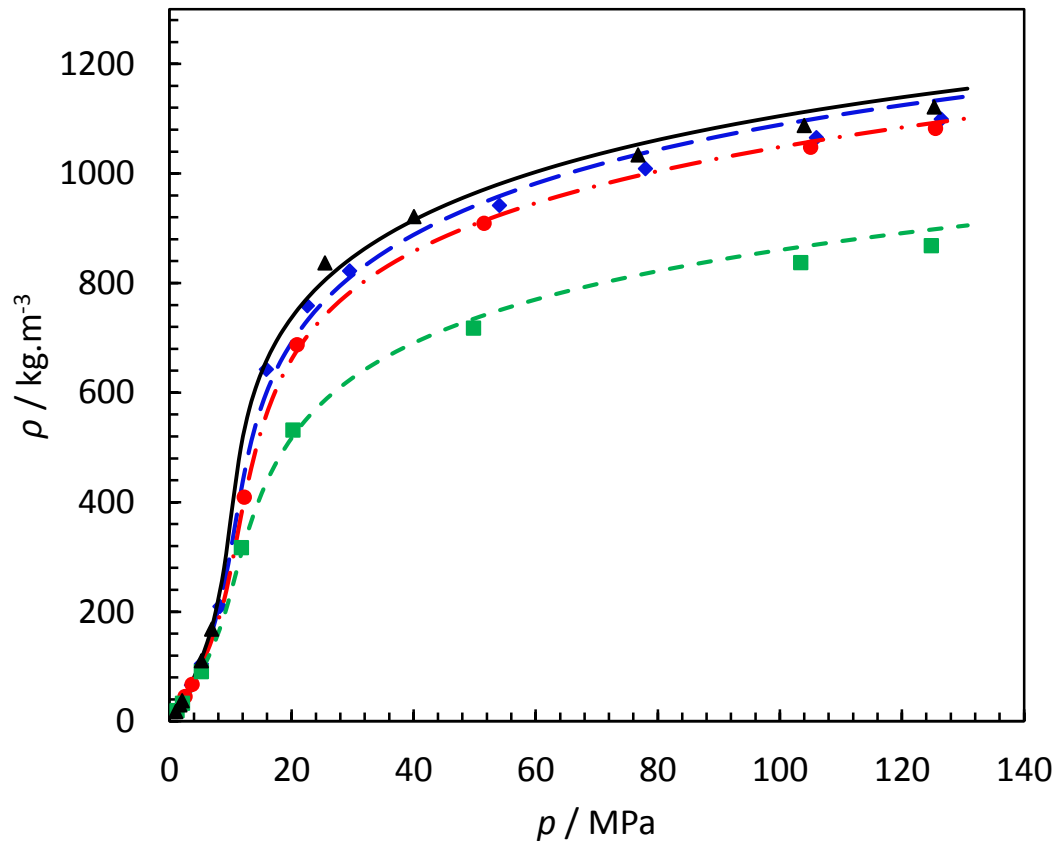


Figure 12. Effect of impurities on the density of pure CO₂ at supercritical area at the temperature of 323.15 K

Pure CO₂: (▲/—), MIX 1: (◆/—), MIX 2: (●/—) and MIX 3: (■/—)

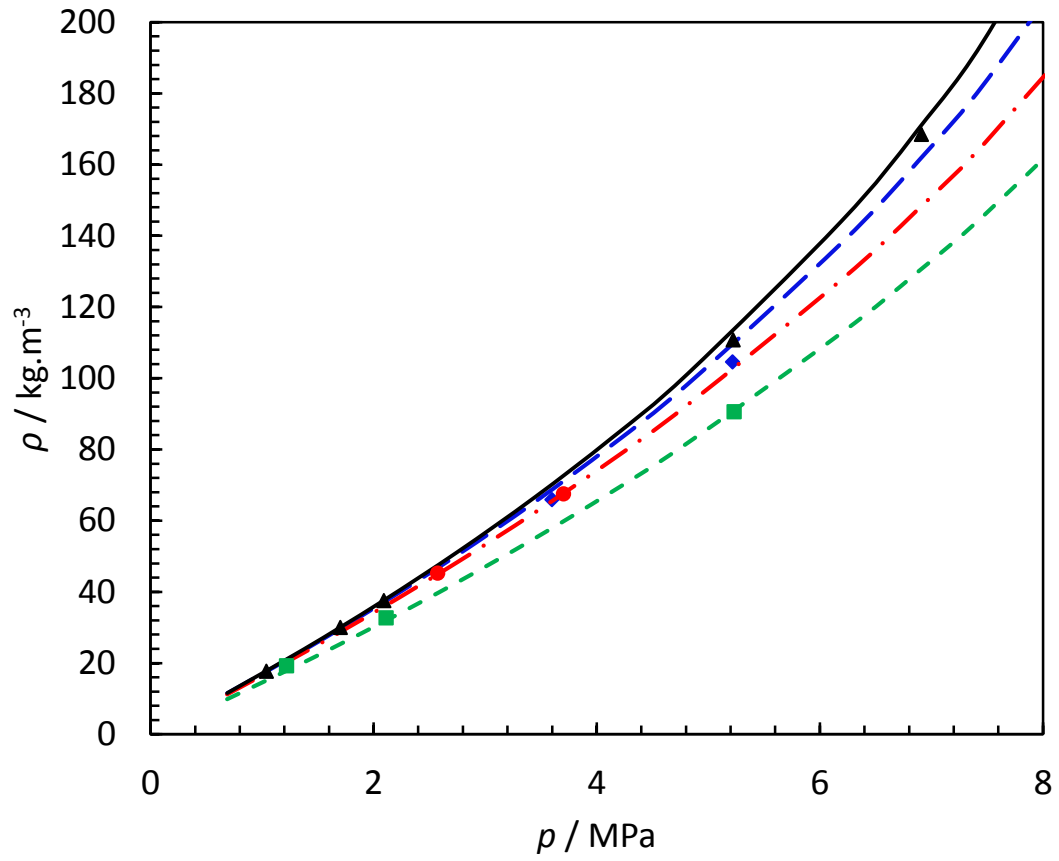


Figure 13. Effect of impurities on the density of pure CO₂ at supercritical area at the temperature of 323.15 K at low pressures

Pure CO₂: (▲/—), MIX 1: (◆/—), MIX 2: (●/—) and MIX 3: (■/—)

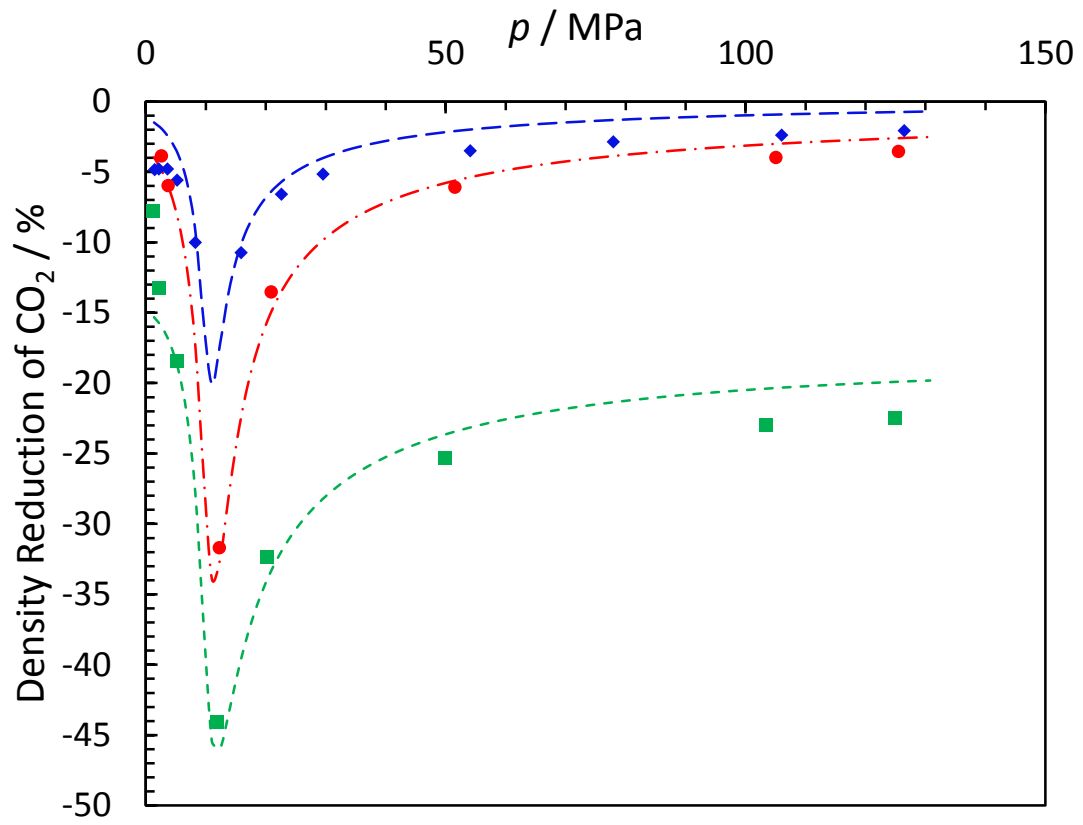


Figure 14. Density reduction of pure CO₂ at supercritical area, temperature 323.15 K (50 °C)
 MIX 1: (♦/—), MIX 2: (●/—) and MIX 3: (■/---)

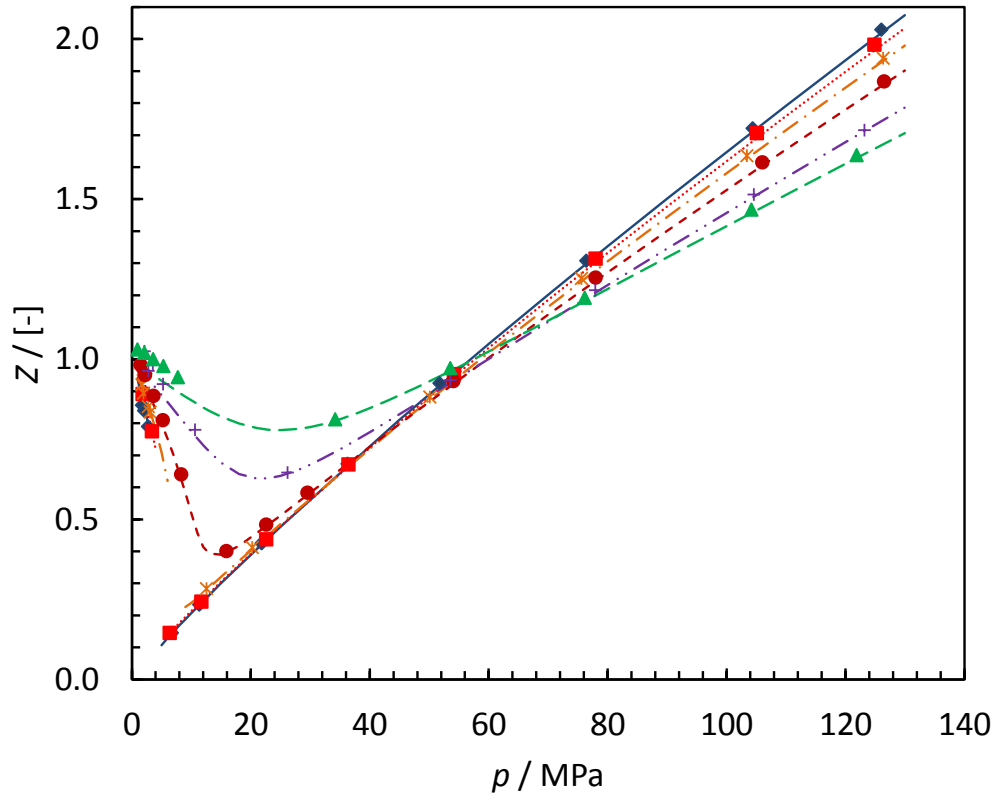


Figure 15. Compressibility factor for MIX 1 at different pressure ranges
 Experimental / modeling (GERG) results: (\blacklozenge / —) at 273.15 K, (\blacksquare / \dots) at 283.15 K, (\times / ---) at 298.15 K, (\bullet / $\text{-}\cdot\text{-}$) at 323.15 K, ($+$ / $\text{-}\cdot\cdot\text{-}$) at 373.15 K and (\blacktriangle / ---) at 423.15 K

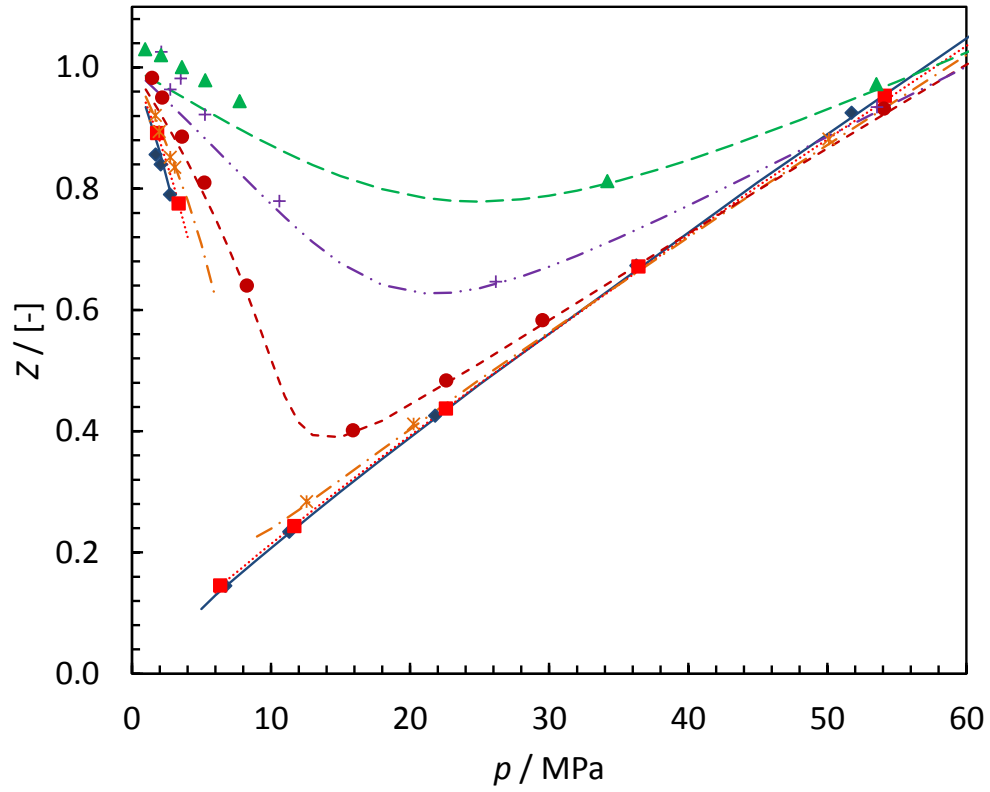


Figure 16. Compressibility factor for MIX 1 at lower pressure ranges
 Experimental / modeling (GERG) results: (\diamond /—) at 273.15 K, (\blacksquare /—) at 283.15 K, (\times /—) at 298.15 K, (\bullet /—) at 323.15 K, ($+$ /—) at 373.15 K and (\blacktriangle /—) at 423.15 K

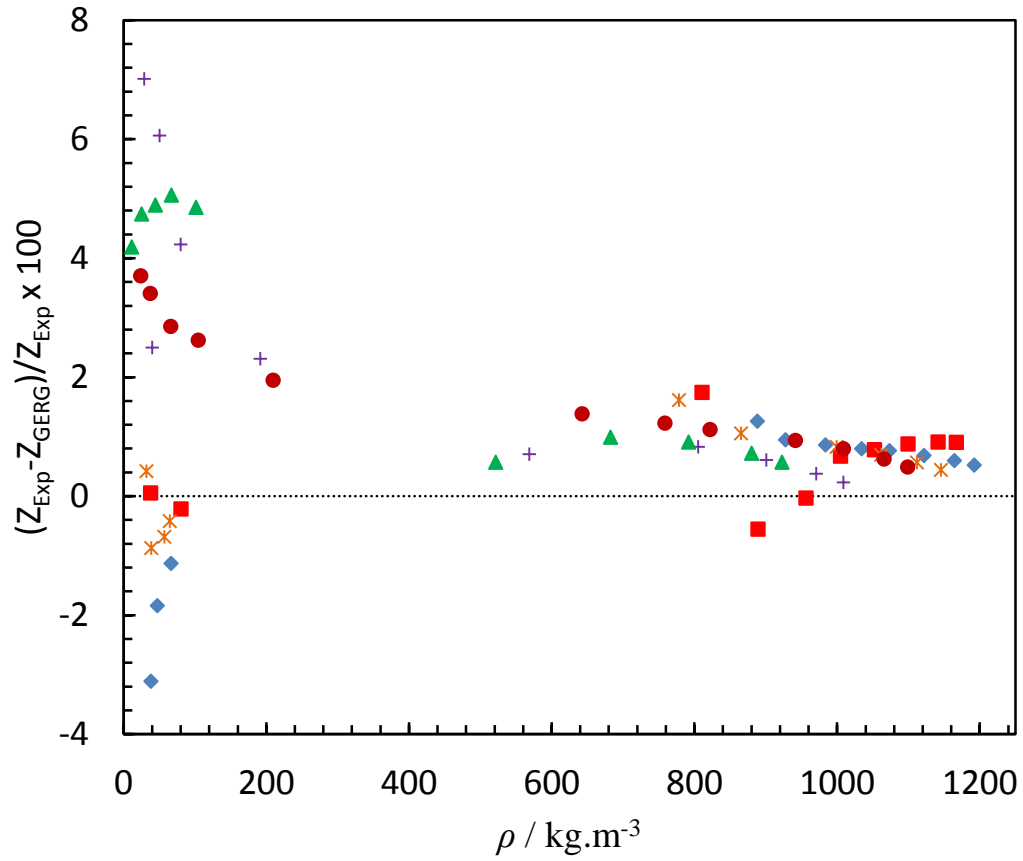


Figure 17. Deviation of compressibility factors (Z) from GERG EoS for MIX 1 at different isotherms, (\blacklozenge) at 273.15 K, (\blacksquare) at 283.15 K, (\times) at 298.15 K, (\bullet) at 323.15 K, ($+$) at 373.15 K and (\blacktriangle) at 423.15 K

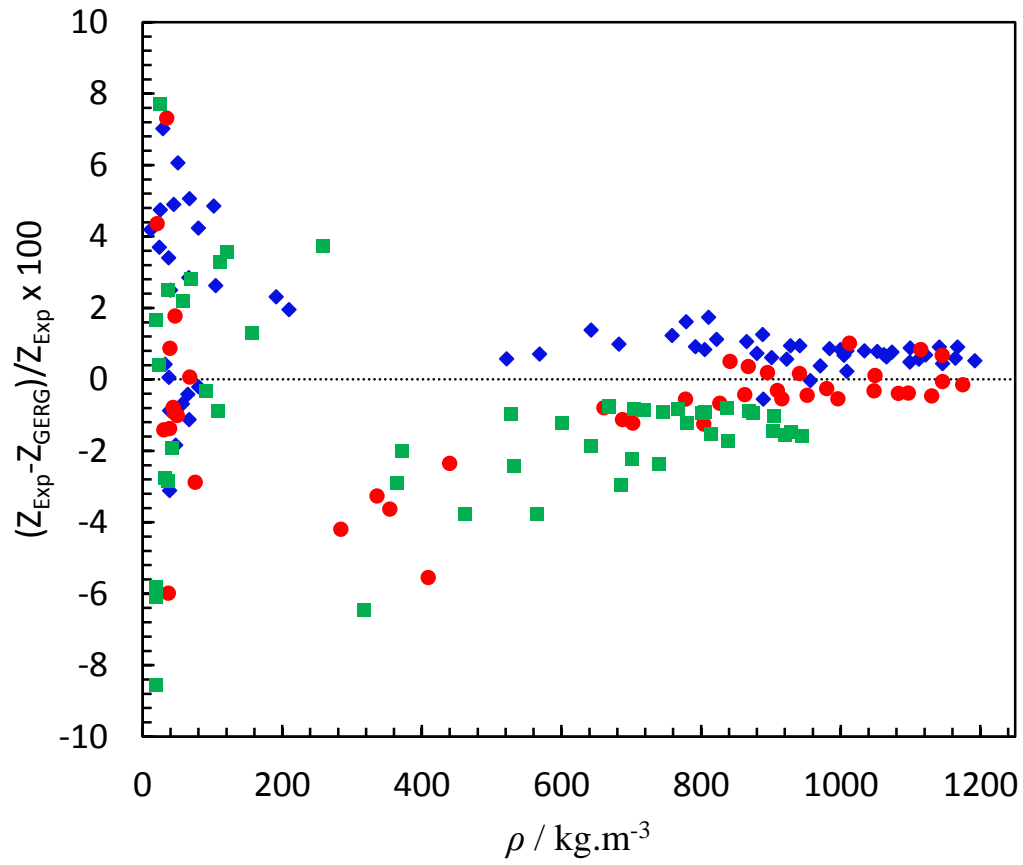


Figure 18. Deviation of compressibility factors (Z) from GERG EoS for investigated mixtures, MIX 1: (♦), MIX 2: (●) and MIX 3: (■)

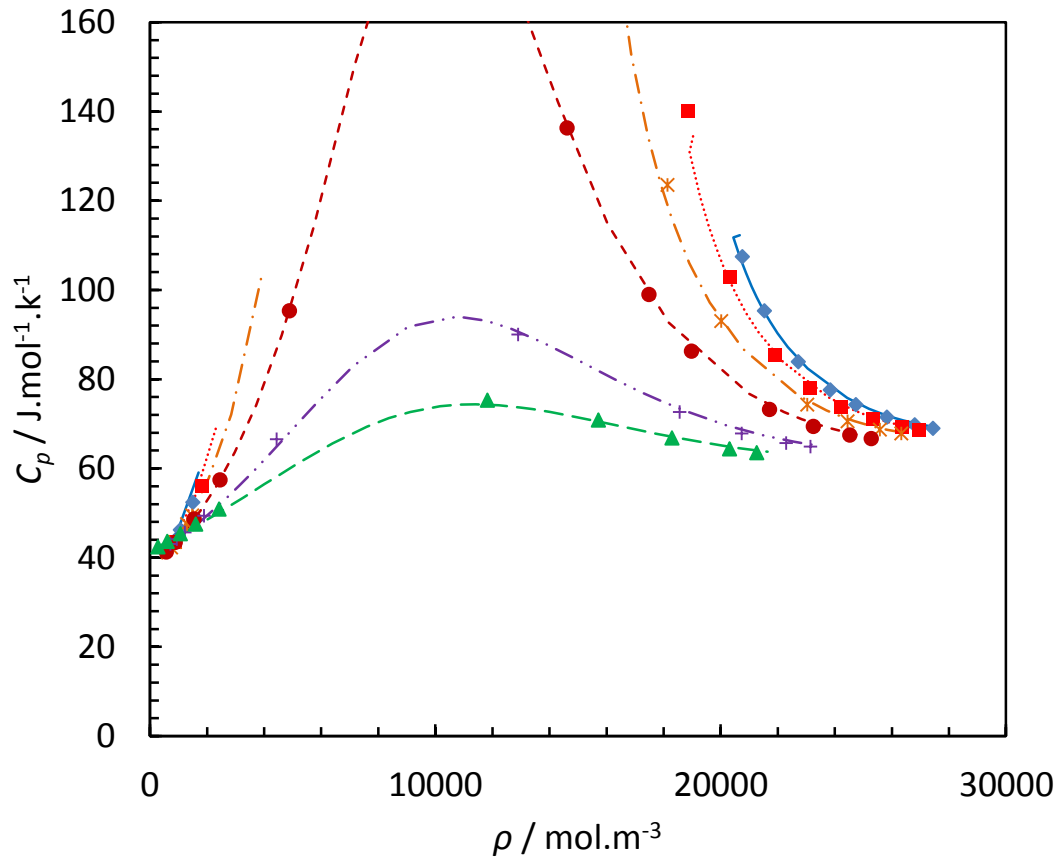


Figure 19. Specific heat capacity (C_p) of MIX 1 at different isotherms
 Experimental / modeling (GERG) results: (\blacklozenge / —) at 273.15 K, (\blacksquare / ---) at 283.15 K, (\times / $\text{-}\cdot\text{-}$) at 298.15 K, (\bullet / $\text{-}\cdot\text{-}$) at 323.15 K, ($+$ / ---) at 373.15 K and (\blacktriangle / —) at 423.15 K

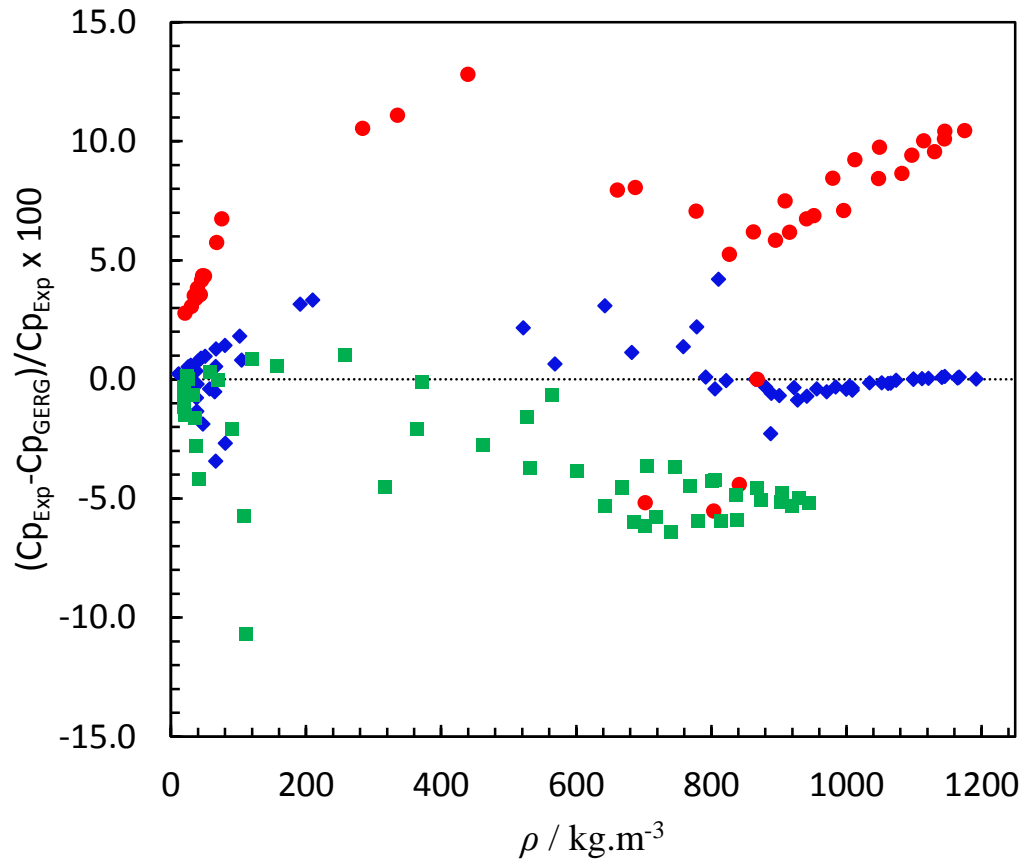


Figure 20. Deviation of specific heat capacity (Cp) from GERG EoS for investigated mixtures, MIX 1: (♦), MIX 2: (●) and MIX 3: (■)

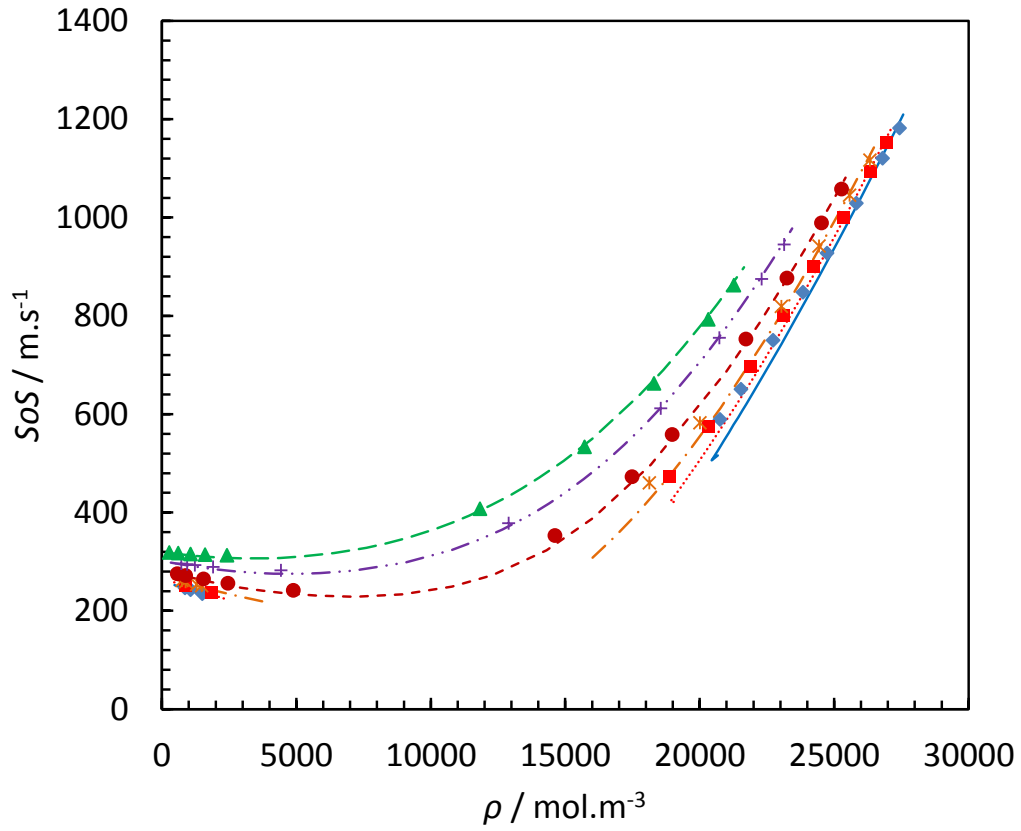


Figure 21. Speed of sound (SoS) of MIX 1 at different isotherms
 Experimental / modeling (GERG) results: (\diamond /—) at 273.15 K, (\square /....) at 283.15 K, (\times /—) at 298.15 K, (\bullet /- - -) at 323.15 K, ($+$ /—) at 373.15 K and (\blacktriangle /—) at 423.15 K

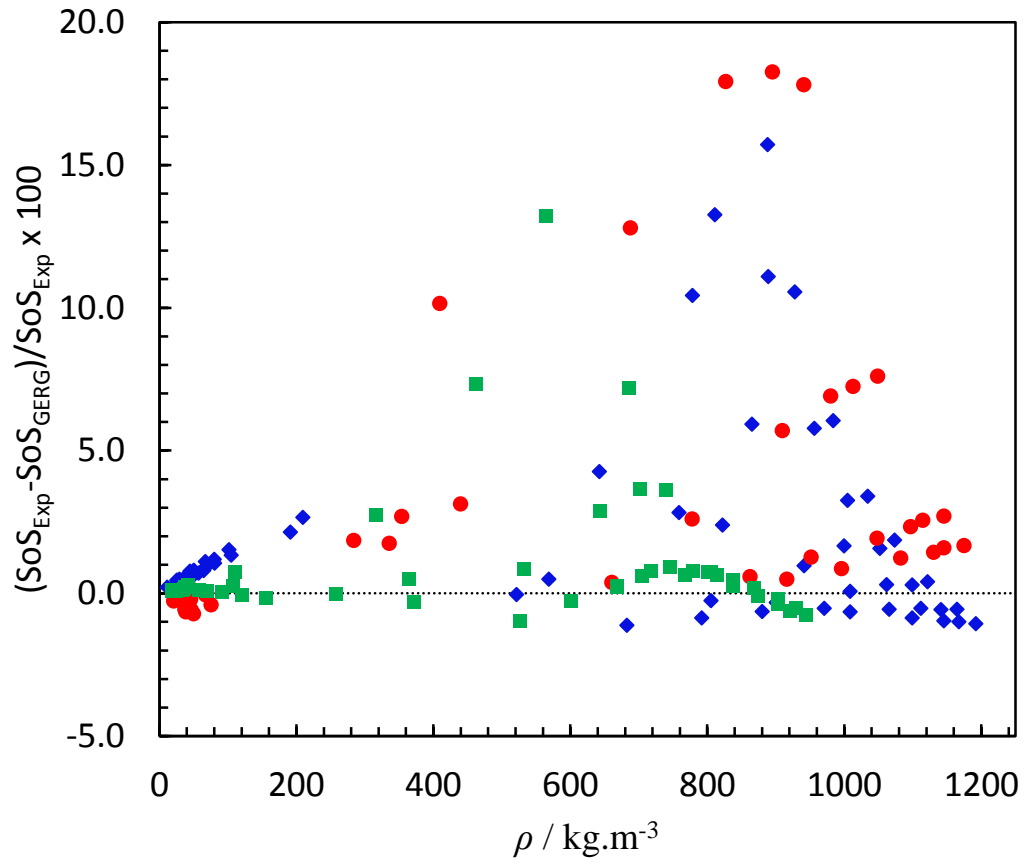


Figure 22. Deviation of speed of sound (SoS) from GERG EoS for investigated mixtures
MIX 1: (♦), MIX 2: (●) and MIX 3: (■)

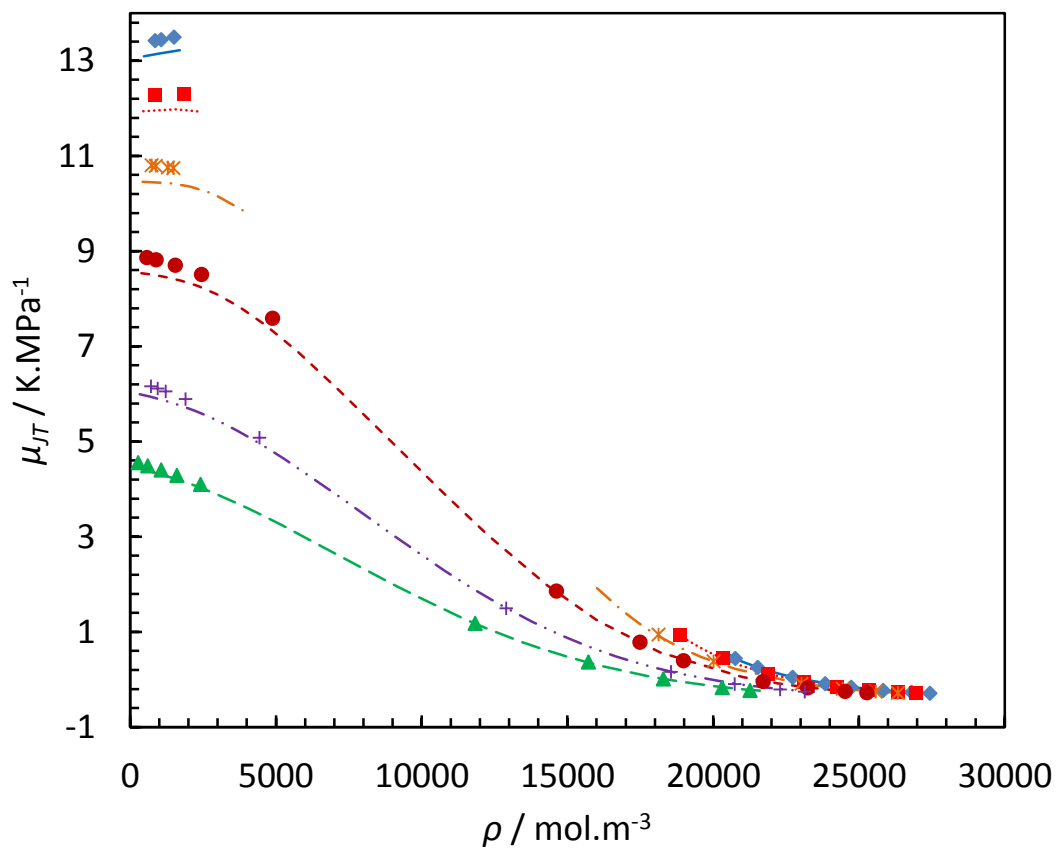


Figure 23. Joule-Thomson Coefficient (μ_{JT}) of MIX 1 at different isotherms
 Experimental / modeling (GERG) results: (\diamond /—) at 273.15 K, (\square /.....) at 283.15 K, (\times /—) at 298.15 K, (\bullet /- - -) at 323.15 K, ($+$ /.....) at 373.15 K and (\blacktriangle /—) at 423.15 K

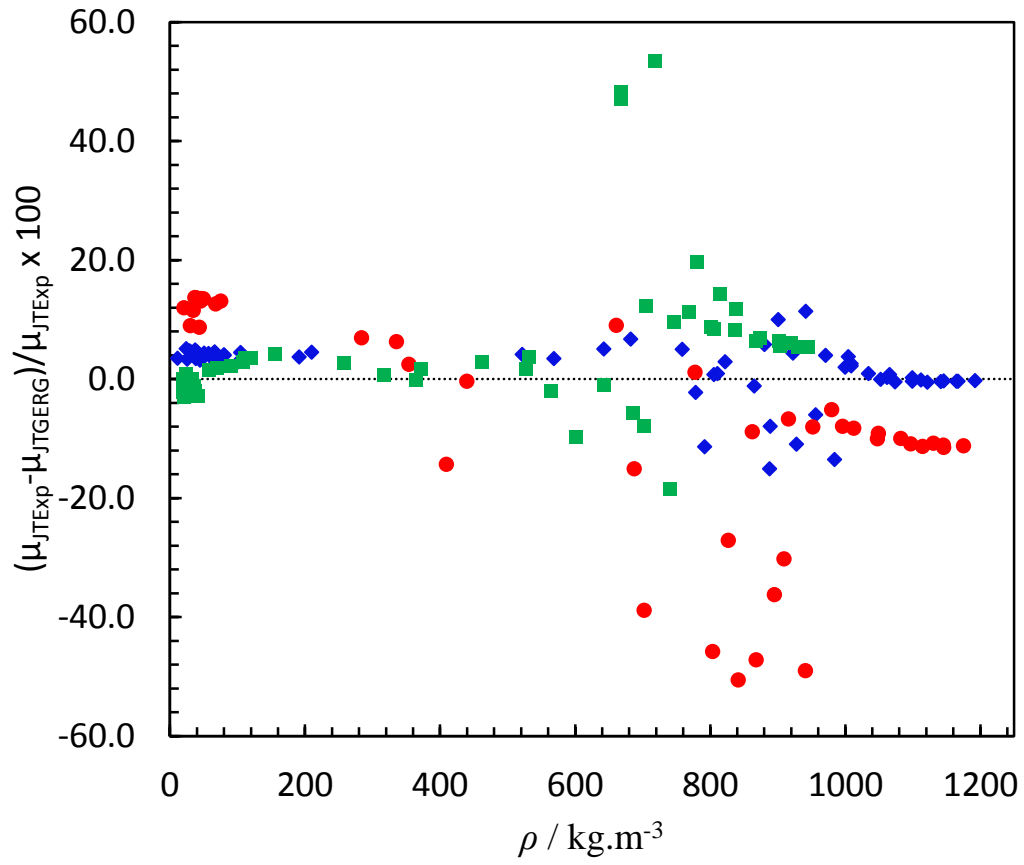


Figure 24. Deviation of Joule-Thomson Coefficient (μ_{JT}) from GERG EoS for investigated mixtures, MIX 1: (\blacklozenge), MIX 2: (\bullet) and MIX 3: (\blacksquare)